

Name:  
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Date:



### Review - Unit 3

#### Topics Covered in Unit 3:

Below are the topics we have covered in Unit 3. The topics below can/may all appear on your test. Please review with your notes, textbook, and practice problems.

- Niels Bohr
  - What did Bohr discover? (*Hint: He used a hydrogen spectrum*)
  - How did Bohr's electron model affect the understanding of the electron and the atom?
    - Can electrons move between energy levels? What would cause them to jump/fall between energy levels?
    - What model did Bohr use to describe his energy levels for electrons?
  - How did Bohr's understanding of the electron relate to light?
    - Be able to move between electron energy levels and light (conceptually).
- Planck's Equation (Light Equation/Math) + Light
  - $E = h \cdot \nu$  (Energy = Planck's constant \* Frequency)
    - $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$
    - Frequency units = (1/s or Hz)
  - $c = \lambda \cdot \nu$  (Speed of Light = Wavelength \* Frequency)
    - $c = 3.00 \times 10^8 \text{ m/s}$
    - Wavelength units = (nm or m) (length)
    - $1.0 \text{ nm} = 1.0 \times 10^{-9} \text{ m}$
  - Understand the electromagnetic radiation light spectrum
    - 7 types of light/radiation
- Orbitals
  - Know the 4 types of orbitals (Do not have to be able to draw all of them...but be able to identify them when seen)
    - s - orbitals
      - 1 type
      - Sphere shaped
    - p - orbitals
      - 3 types
      - Dumbbell shaped
    - d - orbitals
      - 5 types
      - Clover shaped
    - f - orbitals

# Bc Science Unit 3 Review Answers

**Zhe Xu**



### Bc Science Unit 3 Review Answers:

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