

Double Replacement Reactions and Solubility

Net Ionic Equations

Introduction

Precipitation reactions, a type of double replacement reaction, are widely used to prepare new compounds and analyze their purity. Precipitation reactions occur when aqueous solutions of ionic compounds are combined, and a new ionic compound that is insoluble in water is produced. The result is the formation of a precipitate, a solid that settles out of the solution. By carrying out a series of possible double replacement reactions and observing which combinations produce precipitates, we should be able to determine some general rules of solubility for ionic compounds in water.

Concepts

- Double replacement reactions
- Molecular equations
- Net ionic equations
- Solubility rules

Background

Double replacement reactions occur when two ionic compounds (or compounds that break apart to form ions in aqueous solution) exchange ions to form new compounds. Double replacement reactions generally occur in one direction only in response to a “driving force,” which provides a reason for the reaction to occur. Two important driving forces for reactions between ions in aqueous solution are formation of a solid (precipitation reaction) and formation of a gas or soluble molecular product (acid–base neutralization reaction). In the absence of a driving force, ionic compounds will remain dissolved in solution, and no chemical reaction will take place among the dissolved ions.

When an ionic compound dissolves in water, the crystalline solid dissociates or separates into its corresponding cations and anions. For example, silver nitrate dissociates into silver cations and nitrate anions (Equation 1).



When solutions containing two ionic compounds are mixed, two new combinations of cations and anions are possible. In some cases, the cation from one compound and the anion from the other compound may combine to form an insoluble product, a precipitate. For example, when sodium bromide is added to a solution of silver nitrate, there are four

different ions present, and two new combinations of cations and anions—sodium nitrate and silver bromide—are possible. Sodium nitrate is soluble in water, and its ions remain dissolved in solution. Silver bromide is insoluble in water, and precipitates from solution as a solid (Equation 2).

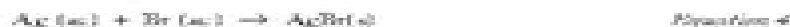


Equation 2 summarizes the double replacement reaction between silver nitrate and sodium bromide. It is called a *molecular equation* because the reactants and products are represented by their molecular formulas. The actual reaction process can be better understood if we rewrite the equation to represent the ionic compounds as they exist in solution, that is, in the form of their dissolved ions.

The *complete ionic equation* for the double replacement reaction of silver nitrate and sodium bromide is shown in Equation 3. Because silver bromide is not soluble, it is best represented by its formula, AgBr.



Notice that some ions (Na^+ and NO_3^-) appear on both sides of the equation in Equation 3. These ions are referred to as *spectator ions* because they do not participate in the overall reaction. Just as in algebra, where a term that appears on both sides of an equation may be “cancelled out,” we can do the same thing with the complete ionic equation shown in Equation 3. Subtracting the spectator ions from both sides gives the *net ionic equation* for the precipitation reaction of silver bromide (Equation 4).



Double replacement reactions may be used to determine some general solubility rules for ionic compounds in water. If a precipitation reaction occurs when solutions of ionic compounds are mixed, then at least one of the two possible products formed by the “ion exchange” reaction must be insoluble in water. In this experiment, various ionic solutions will be mixed one at a time to determine which combinations will form precipitates. Knowing the ions that are present will make it possible to predict the solubility patterns of different ionic compounds.

Double Replacement Reactions And Solubility Net Ionic Equations Answers

Mary Millhollon, Richard H. Langley



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