

Question 1

- 1(b)(iii) From the graph, time taken for the reaction to be completed was 3.2 minutes.
- 1(b)(iv) The rate of the reaction was fastest between time = 0 minutes and time = 0.5 minutes.
- 1(b)(v) The rate of reaction is directly proportional to gradient of the curve. The gradient was steepest within the first 0.5 minutes of the reaction so this means that the rate was fastest during that time. At the start of any reaction, concentrations of reagents are highest. This means that there are the highest number of successful collisions per unit time between reagent particles and the rate is fastest.
- 1(b)(vi) The volume of gas produced at 2.5 minutes was 18.5 cm³.
- 1(b)(vii) 22.4 dm³ = 22400 cm³, so one mole of gas has a volume of 22400 cm³ at S.T.P.
1 cm³ of gas contains (1/22400) mol. gas.
18.5 cm³ contain (18.5 × (1/22400)) mol. gas i.e. 8.26×10^{-4} mol. gas, which is therefore the amount of gas produced in 2.5 minutes.
- 1(b)(viii) It is hydrogen gas.
- 1(c)(i) Andrew is investigating the effect of surface area of a solid reagent on the rate of reaction.
- 1(c)(ii) $\text{Mg}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{MgSO}_{4(aq)} + \text{H}_{2(g)}$
- 1(c)(iii) Temperature at which the reaction is carried out can affect the rate of a reaction.

Question 2

- 2(a)(i) Reduction (of a species) is a decrease in (its) oxidation state.
- 2(a)(ii) CuI is overall neutral. In terms of oxidation states Cu + I = 0. I here exists as I⁻ so Cu + (-1) = 0. Therefore Cu = +1. Hence the oxidation state of Cu in CuI is +1.
- 2(a)(iii) An oxidizing agent is a species which causes the oxidation of another, and is reduced in the process. An oxidizing agent gains an electron or electrons during the reaction and experiences a decrease in oxidation number, or contains at least one species which experiences a decrease in oxidation number/state.
- 2(a)(iv) The Cu²⁺ ion is acting as the oxidizing agent because it experiences a decrease in oxidation number from +2 (Cu²⁺) to +1 (Cu⁺ in CuI).
- 2(a)(v) The solution contained an oxidizing agent.
- 2(a)(vi) The brown colour was due to the presence of I_{2(aq)} which was formed by the oxidation of colourless iodide ions.
- 2(a)(vii) $2\text{I}^{-}_{(aq)} \rightarrow \text{I}_{2(aq)} + 2\text{e}^{-}$

Cxc Chemistry Past Paper Questions

SJ Ball



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| continuous | present | past | future |
| perfect | present | past | future |
| continuous | present | past | future |

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