

Study Card: Enthalpy Change

What is Enthalpy?

- Enthalpy is the total energy possessed by a chemical system
- Represented as H and has units of joules (J) or kilojoules (kJ)
- Includes potential energy associated with particle interactions,
- Includes kinetic energy associated with particle movement
- Includes energy required to make room for the system in the surroundings (often referred to as $P \cdot V$)
- A chemical reaction causes an **enthalpy change** (ΔH)
- At a constant pressure, the ΔH equals the heat (Q) exchanged between the system and the surroundings

Enthalpy Level Diagrams

Enthalpy changes in reactions are often represented by enthalpy level diagrams. The enthalpy of the reaction system is displayed as a function of some indicator of the reaction's progress.

Exothermic Reactions

ΔH is negative (<0)

$H_{\text{reactants}} > H_{\text{products}}$

Heat released by system

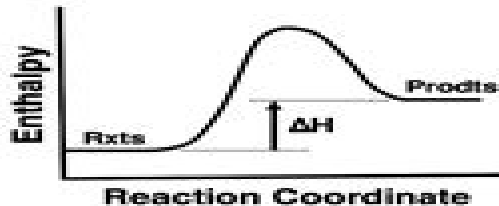


Endothermic Reactions

ΔH is positive (>0)

$H_{\text{reactants}} < H_{\text{products}}$

Heat absorbed by system



System Diagrams

Exothermic Reaction

$$\Delta H < 0$$

$$H_{\text{rxts}} > H_{\text{prodts}}$$



Heat released by system.

Endothermic Reaction

$$\Delta H > 0$$

$$H_{\text{rxts}} < H_{\text{prodts}}$$



Heat absorbed by system.

Thermochemical Equations

Every chemical reaction has a unique enthalpy change or ΔH value. Enthalpy change is a **stoichiometric quantity**; it depends on how much reactants are involved in the reaction. It is often included in the chemical equation for the reaction:



(Translation: 890 kJ of energy are produced by the reaction. $\Delta H = -890 \text{ kJ}$)



(Translation: 572 kJ of energy are required by the reaction. $\Delta H = +572 \text{ kJ}$)

If the coefficients are changed, the heat term changes accordingly. If the reactants and products are reversed, the heat term changes side accordingly:



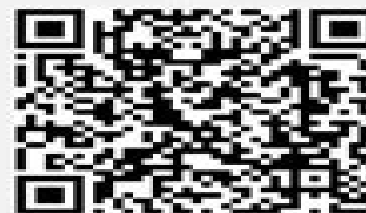
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Causes Of Change Enthalpy Concept Review Answer

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Reviews in Partial Differential Equations, 1980-86, as Printed in Mathematical Reviews, 1988 **Thermodynamic Properties of Aqueous Solutions Organic Substances** V. P. Belousov, Mikhail I. Ust'evich Panov, 1994 Thermodynamic Properties of Aqueous Solutions of Organic Substances discusses the structure of aqueous solutions of organic substances and the intermolecular reactions in them presenting experimental data modern concepts concerning the properties of these solutions and the results of computer simulation The book offers an in depth study of the properties of maximally dilute aqueous solutions of polar and nonpolar organic molecules as well as the specific enthalpies of

mixing The Addendum contains experimental data on the thermodynamic properties of infinitely dilute solutions

CHEMISTRY SILBERBERG,2003 *Biology* Neil A. Campbell,Jane B. Reece,2005 Neil Campbell and Jane Reece s
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Prentice Hall Chemistry ,2000 *Japanese Technical Periodical Index* ,1987

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