

## SALT ANALYSIS

### QUALITATIVE ANALYSIS OF INORGANIC SALT

Date: .....

**AIM:** To determine acid radical and basic radical from the given inorganic salt

PRELIMINARY EXAMINATION			
1	Colour of the sample is noted	The sample is:	Presence of:
		Blue	Cu <sup>2+</sup> ions
		Green	Cu <sup>2+</sup> or Ni <sup>2+</sup> ions
		Dark yellow/Brown	Fe <sup>3+</sup> or Mn(NO <sub>3</sub> ) <sub>2</sub>
		Flesh or light pink	Mn <sup>2+</sup> ions
	Colourless / White	Absence of transition metal cations.	
2	A <u>thick paste</u> of the sample is prepared in con. HCl and is shown to the non-luminous part of a Bunsen flame with the help of a glass rod.	Livid Blue flame	Presence of Pb <sup>2+</sup> ion
		Green-edged blue flame	Presence of Cu <sup>2+</sup> ions.
		Crimson flame	Presence of Sr <sup>2+</sup> ions
		Pale green flame	Presence of Ba <sup>2+</sup> ions
		Brick red flame	Presence of Ca <sup>2+</sup> ion

### ANALYSIS OF ANIONS

	Experiment	Observation	Inference
3	A little of the sample is treated with <b>dil. HCl acid</b>	Colourless, odourless gas (CO <sub>2</sub> ) is evolved with effervescence.	Presence of CO <sub>3</sub> <sup>2-</sup> ion
4	A little of the sample is treated with <b>con. H<sub>2</sub>SO<sub>4</sub></b>	White fuming, pungent smelling gas (HCl) produces dense white fumes when mouth of NH <sub>4</sub> OH bottle is shown near the mouth of the test tube. <b>Note:</b> Warm/Heat if necessary	Presence of Cl <sup>-</sup> ion
5	A pinch of salt is heated with a few drops of conc. H <sub>2</sub> SO <sub>4</sub> for 10-15 seconds.	Reddish brown gas (NO <sub>2</sub> ) on <b>boiling</b> the mixture with copper turnings and a few drops of conc. H <sub>2</sub> SO <sub>4</sub>	Presence of NO <sub>3</sub> <sup>-</sup> ion.
CONFIRMATORY TESTS FOR ANIONS			
6	An aqueous solution of the sample is acidified with <b>dil.HCl</b> followed by the addition of <b>BaCl<sub>2</sub></b> solution.	A white precipitate (BaSO <sub>4</sub> ) formed, which is insoluble in con.HCl	Presence of SO <sub>4</sub> <sup>2-</sup> ion
7	An aqueous solution of the sample is acidified with dilute acetic acid followed by the addition of lead acetate [(CH <sub>3</sub> -COO) <sub>2</sub> Pb] solution.	A white precipitate (PbSO <sub>4</sub> ) formed, which is soluble in Ammonium acetate [CH <sub>3</sub> COONH <sub>4</sub> ] solution.	SO <sub>4</sub> <sup>2-</sup> ion is confirmed.

# **Cbse Salt Analysis Manual**

**Dr. S. C. Rastogi , Er. Meera Goyal**



## **Cbse Salt Analysis Manual:**

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create this wonderful book      **Practical/Laboratory Manual Chemistry Class - XI** Er. Meera Goyal, 2021-05-29

**1 Basic Laboratory Techniques**

- 1 To cut a glass tube or glass rod
- 2 To bend the glass rod at an angle
- 3 To draw a glass jet from a glass tube
- 4 To bore a cork and fit a glass tube into it

**Viva Voce**

**2 Characterisation and Purification of Chemical Substances**

- 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique
- 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method
- 3 To prepare crystals of pure potash alum  $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$  from the given impure sample
- 4 To prepare the pure crystals of copper sulphate from the given crude sample
- 5 To prepare pure crystals of benzoic acid from a given impure sample

**Viva Voce**

**3 Measurement of pH Values**

- 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper
- 2 To determine and compare the pH values of solutions of strong acid HCl and weak acid  $CH_3COOH$  of same concentration
- 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper
- 4 To study the pH change by common ion  $CH_3COO^-$  ion in case of weak acid  $CH_3COOH$
- 5 To determine the change in pH value of weak base  $NH_4OH$  in presence of a common ion  $NH_4^+$

**Viva Voce**

**4 Chemical Equilibrium**

- 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions
- 2 To study the shift in equilibrium between  $Co(H_2O)_6^{2+}$  and  $Cl^-$  ions by changing the concentrations of either of the ions

**Viva Voce**

**5 Quantitative Analysis**

- 1 To prepare M 10 oxalic acid solution by direct weighing method
- 2 To prepare M 10 solution of sodium carbonate by direct weighing method
- 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid
- 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution

**Viva Voce**

**6 Qualitative Analysis**

**Analysis of Anions**

**Analysis of Cations**

**Viva Voce**

**7 Detection of Elements in Organic Compounds**

- 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne's test
- 2 To detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne's test

**Viva Voce**

**INVESTIGATORY PROJECTS**

**1 Checking of Bacterial Contamination in Water**

- 1 To check the bacterial contamination in drinking water by testing sulphide ions

**Viva Voce**

**2 Methods of Water Purification**

- 1 To purify water from suspended impurities by using sedimentation
- 2 To purify water by boiling
- 3 To purify water by distillation method
- 4 To purify water by reverse osmosis technique
- 5 To purify water by GAC method
- 6 To purify water by bleach treatment
- 7 To purify water by oxidising agent
- 8 To purify water by ozone treatment method

**Viva Voce**

**3 Water Analysis**

- 1 To test the hardness of different water samples

**Viva Voce**

**4 Foaming Capacity of Various Soaps**

- 1 To compare the foaming capacity of different washing soaps
- 2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap

**Viva Voce**

**5 Tea Analysis**

- 1 To study the acidity of different samples of tea leaves tea by using pH paper

**Viva Voce**

**6 Analysis of Fruits and Vegetable Juices**

- 1 To analysis the fruit and vegetable juices for the constituent present in them

**Viva Voce**

**7 Rate of Evaporation**

- 1 To study the rate of evaporation of different liquids

IViva Voce 8 Effect of Acids and Bases on Tensile Strength of Fibres 1 To compare the tensile strength of natural fibres and synthetic fibres 2 To study the effect of acids and bases on tensile strength of different fibres Viva Voce

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**Comprehensive Practical Chemistry XII** Dr. N . K. Verma, B. K. Vermani, Dr. Neera Verma, K . K. Rehani, 2011-11

*Salt Analysis Chart* Sibaji SARKAR, 2017-05-27

Salt analysis chart the present book is an attempt and an outcome of the author's experience of teaching the practical of inorganic chemistry subject for more than ten years which will immensely benefit the college and school students of the subject Primary aim to formulate this book is to reflect comprehensively the various aspects of inorganic chemistry and to acquaint students with sound knowledge of performing practical in inorganic chemistry which will be useful to them during the regular internal as well as external examination The book covers the basics of performing salt analysis in a new and exciting way a way that makes it easier for students to learn and a way that will be more interesting for teachers to teach I hope that both the students and teachers of science stream will receive this book favorably as an effective text book I sincerely hope that the book will go a long way to satisfy the long felt need of students for a friendly book on practical in salt analysis I would highly appreciate receiving any input from users which might be of valuable to students at large The author will warmly welcome suggestion from reader at this book and necessary arrangement would be implemented to upgrade quality of book

**Methods for Salt Analysis** Mr. Ajeet Singh, 2018-05-14

Detailed study of qualitative and quantitative methods for identifying and analyzing salts including classical wet chemistry modern instrumental techniques and their industrial applications

**Chemical Student's Manual for the Lecture Room and Laboratory. Embracing non-metals; the Analysis of Simple Salts, Etc** Herbert L. BUCKERIDGE, 1887

Handbook of Water Analysis Leo M.L. Nollet, Leen S. P. De Gelder, 2013-07-29

Extensively revised and updated Handbook of Water Analysis Third Edition provides current analytical techniques for detecting various compounds in water samples Maintaining the detailed and accessible style of the previous editions this third edition demonstrates water sampling and preservation methods by enumerating different ways to measure c

**LABORATORY MANUAL FOR SALT MIXING TEST IN ROD BUNDLES.** , 1978

**Complex Salts** William Thomas, 1924

**Manual of**

**Chemical Analysis ...** C. Greville Williams, 1858      *Laboratory Manual for the Course in Advanced Quantitative Analysis*  
Hobart Hurd Willard, 1920      **Laboratory Manual for Salt-mixing Test in 37- and 217-pin Bundles. [LMFBR].** , 1980

This laboratory manual deals with the procedure employed during salt tracer experiments used in evaluating the hydraulic characteristics of a rod bundle. A description of the standard equipment used is given together with the details of manufacture of probes used for detecting the salt concentration. Details of the bundle construction have been excluded as they are available in the reference cited. An attempt has been made to point out potential trouble areas and procedures.

*Single-salt Analysis* Brian Piers Lascelles, 1896      **Handbook of Pharmaceutical Salts Properties, Selection, and Use** P. Heinrich Stahl, 2008-08-04. This comprehensive up to date guide and information source is an instructive companion for all scientists involved in research and development of drugs and in particular of pharmaceutical dosage forms. The editors have taken care to address every conceivable aspect of the preparation of pharmaceutical salts and present the necessary theoretical foundations as well as a wealth of detailed practical experience in the choice of pharmaceutically active salts. Altogether the contributions reflect the multidisciplinary nature of the science involved in selection of suitable salt forms for new drug products.

**Comprehensive Experimental Chemistry** V. K. Ahluwalia, Sudha Raghav, 1997. This book has been especially written for Class XII students under 10+2 pattern of education according to the syllabi prescribed by the CBSE and other states boards. This book will help the students in acquiring correct skills in practicals and various techniques of all laboratory experiments. Salient features: An introduction to the book is given. This describes the laboratory apparatus and instructions and precautions for working in the laboratory. Simple language and lucid style. Adequate number of illustrations to explain and to clarify the use of various apparatus used in the laboratory. Theoretical aspects of each equipment have been discussed along with experiments. In volumetric analysis both the normality and molarity concepts are made clear. In quantitative analysis inorganic and organic various tests have been given in a systematic way. Specimen recordings of experiments are given to help the students to record on their notebooks. Viva voce questions have been included in each chapter. A fairly large number of investigatory projects covering various topics are given. Selection of projects is carefully made which can be easily performed in school laboratory. An appendix describing various chemical hobbies is given which will be extremely helpful to the students for the development of chemical hobbies. Understanding the basic principles involved and the chemistry of various hobbies. An appendix describing some typical chemical exhibits is also given. This will help the students to participate in the science fairs organized by various agencies. These experiments will cultivate interest among the students for learning chemistry. An appendix each for the solubility of various salts, atomic weights, preparation of various reagents, indicator papers and the first aid to be administered in case of accidents is given. The syllabi prescribed for Class XII students under 10+2 pattern along with distribution of marks is also given.

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