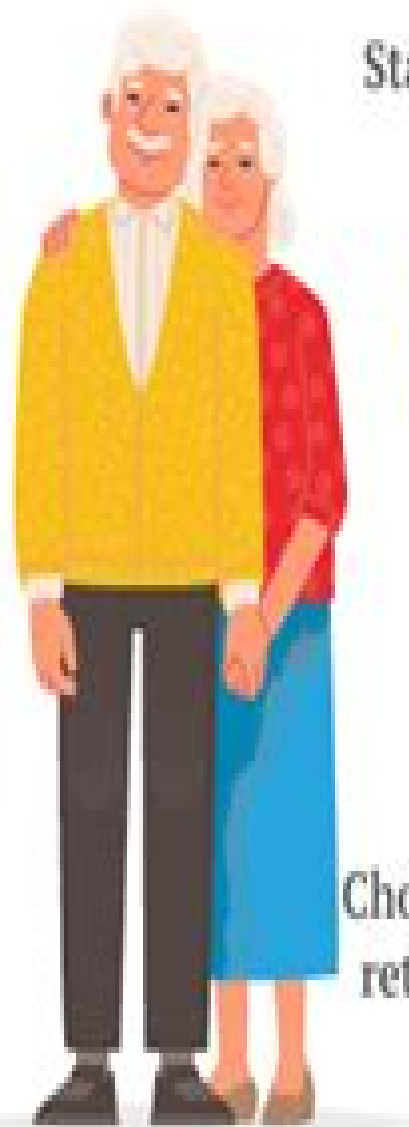


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web unlike strong acids or bases the shape of the titration curve for a weak acid or base depends on the pK_a or pK_b of the weak acid or base being titrated a solution pH as a function of the volume of 1.00 M NaOH added to 10.00 mL of 1.00 M solutions of weak acids with the indicated pK_a values

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web titration curves source openstax chemistry atoms first 2e strong acid strong base titrations when titrating a strong acid such as hydrochloric acid with sodium hydroxide you are reacting the HCl directly with NaOH $HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

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Volume of NaOH added (mL)	pH of the solution
4.6	4.6
4.6	5.9
5.9	5.9

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