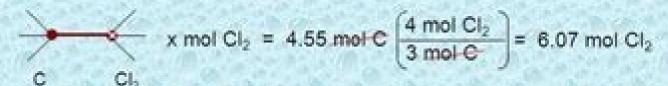
Stoichiometry

$$2 \text{ TiO}_2 + 4 \text{ Cl}_2 + 3 \text{ C} \longrightarrow \text{CO}_2 + 2 \text{ CO} + 2 \text{ TiCl}_4$$
19:5gg x mol 4.55 mol x molecules

How many moles of chlorine will react with 4.55 moles of carbon?



How many grams of titanium (IV) oxide will react with 4.55 moles of carbon?

How many molecules of TiCl4 will react with 115 g TiO2?

x molecules TiCl₄ = 115 g TiO₂
$$\frac{1 \text{ mol TiO}_2}{80 \text{ g TiO}_2} \frac{2 \text{ mol TiCl}_4}{2 \text{ mol TiO}_2} \frac{6.02 \times 10^{23} \text{ molecules TiCl}_4}{1 \text{ mol TiCl}_4}$$



= 8.66x10²³ molecules TiCl₄

Basic Stoichiometry Pogil Answer Key

David Baud

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