

Date Performed: 9/18/06
Date Submitted: 9/25/06
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Determination of a Chemical Formula

Objective:

To determine the percentage composition and simplest formula for an oxide of copper by its reduction with methane gas (CH_4) at approximately 500°C . This experiment is a study of stoichiometry.

Experimental Data:

Mass of Test Tube	$17.726 \pm .001 \text{ g}$
Mass of Test Tube and Cu_2O	$19.705 \pm .001 \text{ g}$
Mass of Cu_2O	$1.979 \pm .001 \text{ g}$
Mass of Test Tube and Cu	$19.313 \pm .001 \text{ g}$
Mass of Cu	$1.587 \pm .001 \text{ g}$
Mass of O	$0.392 \pm .001 \text{ g}$

Sample Calculation:

Mass percentage calculation:

$$\frac{\text{Mass of Atom}}{\text{Mass of Compound}} \times 100 = \frac{1.587 \pm .001 \text{ g of Cu}}{1.979 \pm .001 \text{ g of Cu}_2\text{O}} \times 100 = 79.73\%$$

Calculation of a Chemical Formula:

Assume a 100 g compound and calculate individual mass using mass percentage.

$$79.73 \text{ g of Cu} \times \frac{1 \text{ mol of Cu}}{63.55 \text{ g per mol}} = 1.255 \text{ mol of Cu}$$

$$20.07 \text{ g of O} \times \frac{1 \text{ mol of O}}{16.00 \text{ g per mol}} = 1.254 \text{ mol of O}$$

Results and Conclusions:

Mass of Test Tube	$17.726 \pm .001 \text{ g}$
Mass of Test Tube and Cu_2O	$19.705 \pm .001 \text{ g}$
Mass of Cu_2O	$1.979 \pm .001 \text{ g}$
Mass of Test Tube and Cu	$19.313 \pm .001 \text{ g}$
Mass of Cu	$1.587 \pm .001 \text{ g}$
Mass of O	$0.392 \pm .001 \text{ g}$
Mass Percentage of Cu	79.73%
Mass Percentage of O	20.07%

Cu and O have a one to one mole ratio; therefore the empirical formula is CuO .

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