

## DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Solve the problems below

You may upload a picture of your work on paper, if you prefer...

1. The empirical formula of a compound is  $\text{NO}_2$ . Its molecular mass is 92 g/mol. What is its molecular formula?

2. The empirical formula of a compound is  $\text{CH}_2$ . Its molecular mass is 70 g/mol. What is its molecular formula?

3. A compound is found to be 40.0% carbon, 6.7% hydrogen, and 53.5% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

4. A compound is 64.9% carbon, 13.5% hydrogen, and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

5. A compound is 54.5% carbon, 9.1% hydrogen, and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

# Determining Molecular Formulas True Formulas Answer Keys

**Ralph H. Petrucci, William S. Harwood**



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