

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Derive the molecular formula:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Given: empirical formula =

Mass

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Given: CH_2 is a 14 g/mol. Its molecular mass is

Mass

3. A compound is found to contain 60.0% carbon, 4.48% hydrogen and 35.5% oxygen. Its molecular mass is 106 g/mol. What is its molecular formula?

Given: 60.0% carbon, 4.48% hydrogen, 35.5% oxygen

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Mass

4. A compound is found to contain 60.0% carbon, 4.48% hydrogen and 35.5% oxygen. Its molecular mass is 106 g/mol. What is its molecular formula?

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Mass

Determining Molecular Formulas True Formulas Answer Key

**George A. Williams, Richard
Barnes, Ray Doerhoff, Max Bolen**

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The book delves into Determining Molecular Formulas True Formulas Answer Key. Determining Molecular Formulas True Formulas Answer Key is a crucial topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Determining Molecular Formulas True Formulas Answer Key, encompassing both the fundamentals and more intricate discussions.

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- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Determining Molecular Formulas True Formulas Answer Key.

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