



# Bmw 118d M Sport Owners Manual

**J Dewey**



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 via and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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eg 1  $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o l}$  for which  $\Delta h = 57\text{ kJ mol}^{-1}$  eg 2  $\text{h}_2\text{so}_4 \text{ aq} + \text{koh aq} \rightarrow \text{ks}_2\text{so}_4 \text{ aq} + \text{h}_2\text{o l}$  for which  $\Delta h = 57\text{ kJ mol}^{-1}$

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