

B = Bonding & Molecular Structure

Table B.1 • Some Average Single and Multiple Bond Energies (kJ/mol)

	H	Li	Na	K	Rb	Cs	F	Cl	Br	I	At
H-H	436						485	242	200	167	
Li-H		103									
Na-H			91								
K-H				86							
Rb-H					82						
Cs-H						76					
F-F							158	242	193	171	
Cl-Cl								242	203	177	
Br-Br									193	163	
I-I										151	
At-At											122
C-C							347	336	276	234	
C-H							413	411	399	377	
C-Cl							328	327	309	285	
C-Br							276	275	248	232	
C-I							238	234	213	193	
C-At							220	216	193	173	
C≡C							837	837	812	799	
C≡N							891	891	860	840	
N-N							160	158	146	141	
N-O							469	444	417	401	
O-O							146	142	138	133	
O=O							498	494	470	460	
S-S							226	226	217	208	
S=O							523	500	485	468	
S=O ₂							548	523	507	490	
S=O ₃							551	531	515	498	
Se-Se							192	192	185	178	
Se=O							342	330	315	305	
Se=O ₂							360	348	333	323	
Te-Te							126	126	121	116	
Te=O							338	326	311	301	
Te=O ₂							354	342	327	317	

Multiple Bonds			
C≡C	837	C≡N	891
C=C	614	C=N	889
C=O	799	C=O	799
C=O	743	C=O	743
C≡O	1072	C≡O	1072
C≡O (gas, $D_{0,0}$)	1078	C≡O (gas, $D_{0,0}$)	1078

1. Write the balanced equation for the combustion of 1 mole of glucose, $C_6H_{12}O_6$, completely in O_2 .



2. Draw the structural formula showing Lewis structures for each of the molecules involved.

B = 1.07 C = 0.76 O = 0.34 H = 0.21 N = 0.30 S = 0.35 P = 0.21



3. Calculate the bond energies of the reactants.

Bond	kJ/mol	kJ/mol
C-H	374	374
C-O	336	336
O-H	463	463

4. Calculate the bond energies of the products.

Bond	kJ/mol	kJ/mol
C=O	799	799
O-H	463	463

Total Bond Energy (Reactants)	1110 kJ	Total Bond Energy (Products)	5478 kJ
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5. Calculate the Heat Energy released. Bond Energy = _____

What is the maximum heat released? -4368 kJ

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Linus Pauling



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Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1989 Designed for use in inorganic physical and quantum chemistry courses this textbook includes numerous questions and problems at the end of each chapter and an Appendix with answers to most of the problems

Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1965

Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973

Structure and Bonding Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes

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Bonding and Molecular Structure Russell S. Drago, 1971

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A Pictorial Approach to Molecular Bonding John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example

Molecules and the Chemical Bond Henry A. Bent, 2011 MOLECULES AND THE CHEMICAL BOND Chemistry Simplified This highly original book by a famous chemistry teacher about general chemistry in a new key may change how teachers teach Atomic Theory The Mole Concept and Avogadro's Constant The Gas Laws Solving Problems in Chemical Stoichiometry The Saturation and Directional Character of Chemical Affinity The Pauli Exclusion Principle Linnett's Double Spin Set Theory Pauling's Rules of Crystal Chemistry The Octet Rule Lewis Structures for O_2 NO CO SO_2 and SO_3 Construction of Bond Diagrams VSEPR Theory Dative Bonding Multicenter Bonding Bonding in Metals pH Calculations The

Periodic Table The Energy Function and the First Law of Thermodynamics The Entropy Function and the Second Law of Thermodynamics How an Inductive Science Advances Molecular Structure and Conformation I. G. Csizmadia, 1982 Good No Highlights No Markup all pages are intact Slight Shelfwear may have the corners slightly dented may have slight color changes slightly damaged spine *The Theory of Molecular Structure and Bonding*, 1980 **Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1967 **Structure of Molecules and the Chemical Bond** И. А. Кивович Сыркин, М. Е. Дикаткина, 1964 **The Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1966 **The Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Carl Pauling, 1993 **The Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1966 The Nature of the Chemical Bond Linus Pauling, 1964 Essentials of Chemical Bonding John Stanton, 2025 Fundamental Concepts Comprehensive coverage of fundamental concepts related to chemical bonding including atomic structure electron configuration and the periodic table Types of Chemical Bonds Explanation of different types of chemical bonds such as covalent bonds ionic bonds metallic bonds and hydrogen bonds including their formation properties and significance in chemical reactions Molecular Structure and Geometry Exploration of molecular structure and geometry including the VSEPR theory Valence Shell Electron Pair Repulsion molecular orbital theory and the relationship between molecular structure and chemical properties Bonding Theories and Models Discussion of various bonding theories and models such as Lewis structures valence bond theory and molecular orbital theory with examples and applications to illustrate their predictive power in explaining chemical behavior Chemical Bonding in Practice Application of chemical bonding principles to real world scenarios including examples from organic chemistry inorganic chemistry materials science and biochemistry demonstrating how an understanding of chemical bonding is essential for understanding and manipulating matter at the molecular level **Molecular Structure and Energetics; Volume 1: Chemical Bonding Models**, 1986 The Determination of Molecular Structure Peter Jaffrey Wheatley, 1962

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