

B = Bonding & Molecular Structure

Table B.3 • Some Average Single- and Multiple-Bond Energies (kJ/mol)

	H	C	N	O	F	Cl	S	Br	I
H-H	436								
C-H		413							
C-C		347							
C=C		614							
C≡C		839							
N-H			391						
N-N			160						
N=O			607						
O-H				467					
O-O				201					
S-H					347				
S-S					226				
Br-Br						224			
I-I							151		
At-At								94	
C-Cl						328			
C-Br						276			
C-I						213			
C-At							201		
C-S							258		
C-Se							226		
C-Te							182		
C-P							200		
C-As							181		
C-Sb							170		
C-Bi							126		
C-Si							318		
C-Ge							186		
C-Sn							147		
C-Pb							126		
C-B							400		
C-Al							224		
C-Ga							186		
C-In							147		
C-Tl							126		
C-Mg							74		
C-Ca							93		
C-Fe							104		
C-Ni							120		
C-Cu							120		
C-Ag							114		
C-Au							100		
C-Hg							93		
C-Zn							104		
C-Co							114		
C-Ni							120		
C-Pd							100		
C-Pl							93		
C-Mn							114		
C-Cr							104		
C-Mo							114		
C-W							100		
C-V							93		
C-Ta							114		
C-Nb							104		
C-Ti							114		
C-Zr							104		
C-Hf							114		
C-U							104		
C-Th							93		
C-La							93		
C-Ba							93		
C-Sr							93		
C-Ca							93		
C-Mg							74		
C-Be							224		
C-Li							93		
C-K							93		
C-Rb							93		
C-Cs							93		
C-Fr							93		
C-Ra							93		
C-Act							93		
C-Th							93		
C-U							104		
C-Np							104		
C-Pu							104		
C-Am							104		
C-Cm							104		
C-Bk							104		
C-Cf							104		
C-Er							104		
C-Tm							104		
C-Yb							104		
C-Lu							104		
C-Hf							114		
C-Ta							114		
C-W							100		
C-Re							100		
C-Mo							114		
C-Tc							104		
C-Co							114		
C-Ni							120		
C-Cu							120		
C-Zn							104		
C-Ga							186		
C-Ge							186		
C-As							186		
C-Se							226		
C-Br							276		
C-I							213		
C-At							201		
C-Bi							126		
C-Pb							126		
C-Tl							126		
C-Po							126		
C-Fl							126		
C-Mg							74		
C-Ca							93		
C-Fe							104		
C-Ni							120		
C-Cu							120		
C-Ag							114		
C-Au							100		

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Roger L. DeKock, Harry B. Gray



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Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973 *Structure and Bonding* Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes *Bonding and Molecular Structure* Russell S. Drago, 1971 **A Pictorial**

Approach to Molecular Bonding John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH for example **The Shape and Structure of Molecules** Charles Alfred Coulson, 1973

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