

CHAPTER 2

Electrochemistry

SUMMARY

1. ELECTROLYTIC CELL

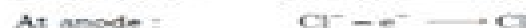
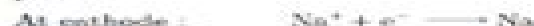
It is a device in which electricity from outside source is used to carry out a reaction (electrolysis). The device or cell used to carry out electrolysis is called electrolytic cell.

2. ELECTROCHEMICAL CELL OR GALVANIC CELL

It is a device in which indirect redox reaction takes place and decrease in energy during the reaction appears in the form of electricity.

3. ELECTROLYSIS OF MOLTEN NaCl

In molten state, NaCl consists of Na^+ consist Cl^- ions. On passing electricity, following reactions takes place :



The over all reaction :



4. ELECTROLYSIS OF ACIDIFIED WATER

It consists of H^+ , OH^- , SO_4^{2-} ions.

On passing electricity, following reactions takes place :



The overall reaction is :



5. ELECTROLYSIS OF CONCENTRATED AQUEOUS NaCl

It consists of Na^+ , H^+ , Cl^- and OH^- ions.

On passing electricity, following reactions takes place :



The overall reaction is :



The solution becomes basic in nature due to the formation of NaOH.

6. ELECTROLYSIS OF AQUEOUS CuSO_4

It consists of Cu^{2+} , H^+ , SO_4^{2-} and OH^- ions.

On passing electricity, following reaction takes place :



The overall reaction is :



The solution becomes acidic in nature due to the formation of H_2SO_4 .

7. ELECTROLYSIS OF CuSO_4 USING Cu ELECTRODES

It consists of Cu^{2+} , H^+ , SO_4^{2-} and OH^- ions.

On passing electricity, following reactions takes place :



8. FARADAY'S LAW OF ELECTROLYSIS

(a) The amount of substance deposited at any electrode during electrolysis is directly proportional to the quantity of electricity passing i.e.,

$$W \propto Q$$

$$W = zQ$$

$$\text{or} \quad W = zIt$$

Where, z = electrochemical equivalent,

W = amount of the substance deposited (g),

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