

MODULE C : SOLIDS, LIQUIDS & GASES

Kinetic Molecular Theory & Phases of Matter

- The different phases of matter can be described based on the Kinetic Molecular Theory

a) According to kinetic theory, an increase in temperature will cause the average kinetic energy of the particles in the solid, liquid, and gas to increase.

↑ temperature ↑ average kinetic energy

b) When a solid is heated, the temperature is raised and the solid changes to a liquid at its melting point. If the heating is continued, the liquid will change into a gas at its boiling point.

c) The heat energy is changed into the equivalent amount of kinetic energy. Thus, the kinetic energy content of the particles increases in the order:

solid < liquid < gas

- Kinetic Molecular Theory

- 1) Matter is made up of particles
- 2) The particles are always moving

- Kinetic energy

→ energy of movement

- Kinetic Molecular Theory for Gases

a) A gas consists of small particles (atoms or molecules) which have negligible volume (V of gas molecule ≈ 0). This implies that:

$$V_{\text{container}} = V_{\text{molecule}} + V_{\text{free space}} \approx V_{\text{free space}}$$

and that there is plenty of space for compression.

b) Gas particles are separated relatively far from one another.

c) Gas particles behave independently from one another. Thus, inter-particle forces of attraction are negligible.

d) The gas particles are in rapid and continuous motion.

e) In the gaseous state, the gas particles have vibrational, rotational, and translational motion.

1) **Vibrational motion** - a particle vibrates about a fixed point

2) **Rotational motion** - a particle spins about a fixed point

3) **Translational motion** - a particle moves from one place to another

f) Pressure of a gas comes from molecules colliding with the walls of the container.

Such collisions are perfectly elastic.

g) Average kinetic energy depends only on the absolute temperature in Kelvin (K)

$$\text{Average Kinetic Energy} = \frac{3}{2} RT, \quad R \rightarrow \text{gas law constant } (0.08206 \frac{\text{L-atm}}{\text{K-mol}})$$

$T \rightarrow \text{temperature (in Kelvin)}$

Chapter Review Part B Concept Review Answers

Paul A. Foerster



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Research Handbook on International Procedural Law Joanna Gomula, Stephan Wittich, Markus Stemeseder, 2024-06-05 This comprehensive Research Handbook provides a detailed exploration of the principles and rules that impact the procedures and operation of international courts and tribunals Within this framework leading experts examine how the evolution of procedural rules and concepts has given rise to a distinct body of rules known as international procedural law *Glencoe Science* McGraw-Hill Staff, 2001-08 **Glencoe Sci Earth Science Chapter 14 Geologic Time Chp Res 513 2002** McGraw-Hill Staff, 2001-08 **Glencoe Sci Earth Science Chapter 21 Our Impact on Water and Air Chp Res 519 02** McGraw-Hill Staff, 2001-08 **Fundamentals of Microbiology** Jeffrey C. Pommerville, 2014-12 Ideal for health science and nursing students Fundamentals of Microbiology Body Systems Edition Third Edition retains the engaging student friendly style and active learning approach for which award winning author and educator Jeffrey Pommerville is known Highly suitable for non science majors the fully revised and updated third edition of this bestselling text contains new pedagogical elements and an established learning design format that improves comprehension and retention and makes learning more enjoyable Unlike other texts in the field Fundamentals of Microbiology Body Systems Edition takes a global perspective on microbiology and infectious disease and supports students in self evaluation and concept absorption Furthermore it includes real life examples to help students understand the significance of a concept and its application in today s world whether to their local community or beyond New information pertinent to nursing and health sciences has been added while many figures and tables have been updated revised and or reorganized for clarity Comprehensive yet accessible the Third Edition is an essential text for non science majors in health science and nursing programs taking an introductory microbiology course Provided by publisher Focus on physical science Charles H. Heimler, 1989 Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems Interactions of Life , *Study Guide to accompany Basic Concepts of Chemistry, 7th Edition* Leo J. Malone, 2003-02-20 Work more effectively and gauge your progress along the way This Study Guide that is designed to accompany Malone s Chemistry 7th Edition includes chapter summaries new terms self tests answers to self tests and solutions to selected problems This easy to read introduction presents chemistry as a living relevant science Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from

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