

Acids and Bases Worksheet 2

1. In the reaction: $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$
the carbonate ion is acting as a(n):
(A) Arrhenius base
(B) Arrhenius acid
(C) Bronsted-Lowry base
(D) Bronsted-Lowry acid
2. Which of the following represents a Bronsted-Lowry conjugate acid-base pair?
(A) SO_3^{2-} and SO_2 (B) CO_3^{2-} and CO (C) H_3O^+ and H_2 (D) NH_4^+ and NH_3
3. Which of the following would not make a good buffering system?
(A) SO_4^{2-} and H_2SO_4
(B) HCO_3^- and H_2CO_3
(C) NH_3 and NH_4^+
(D) CH_3COO^- and CH_3COOH
4. A water solution of which of the following compounds will turn blue litmus red
(A) K_2CO_3 (B) NH_4Cl (C) NaOH (D) NaCl
5. 10.0 ml of a solution of HCl required 12.5 ml of 0.400 M $\text{Ba}(\text{OH})_2$ for complete neutralization. How many moles of HCl were present in the sample?
(A) 5.00×10^{-3} (B) 1.00×10^{-2} (C) 1.00 (D) 2.00
6. If 1.00 L of 1.00 M CH_3COOH is mixed with 0.25 mole of solid NaOH (assume no volume change), what will be the pH of the resulting solution? ($K_a = 1.8 \times 10^{-5}$).
(A) 4.14 (B) 4.27 (C) 4.74 (D) 5.35
7. What is the molarity of CH_3COOH in vinegar containing 4.0% CH_3COOH by mass and having a density of 1.02 g/ml?
(A) 0.50 M (B) 0.68 M (C) 0.75 M (D) 1.36 M
8. If 5.00 ml of 15.4 M HNO_3 is diluted to 250 ml, what is the pH of the resulting solution?
(A) 0.51 (B) 0.76 (C) 1.45 (D) 2.89

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Cathy Duffy



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