

Chapter 13 Homework Answers

Period 1

13.12.

- a) dispersion forces
- b) hydrogen bonding
- c) dispersion forces

13.10 b and c (only)

- b) dipole-dipole
- c) dipole-induced dipole

13.19

The more carbon and hydrogen atoms present, the more soluble the substance is in non-polar oil droplets. Therefore, sodium propanoate is not as effective a soap as sodium stearate with the longer hydrocarbon chain.

13.26 a and b (only)

- a) Charge density is the ratio of the ion's charge to its volume. An ion's charge and size affect its charge density.
- b) $- < + < 2- < 3+$

13.33 a,b and c (only)

- a) I^- has a smaller charge density (larger ion volume) than Br^- .
- b) Ca^{2+} is less than Sc^{3+} , due to its smaller ion charge.
- c) Br^- is less than K^+ , due to its larger ion volume.
- d) Cl^- is less than Si^{4-} , due to its smaller ion charge.
- e) Se^{4+} is less than Al^{3+} , due to its larger ion volume.
- f) ClO_4^- has the smaller charge density due to its smaller ion charge.
- g) Fe^{3+} has the smaller charge density due to its smaller ion charge.
- h) K^+ has the smaller charge density due to its smaller ion charge.

Period 2

13.42



Prepare a mixture of more than 6.4 g KMnO_4 / 100 g H_2O and heat it until the solid completely dissolves. Then carefully cool it, without disturbing it or shaking it, back to 20°C . If no crystals form, you would then have a supersaturated solution.

13.53

No, 21 g solute/kg of solvent would be 21 g solute / 1.021 kg solution.

13.61

- a) Find the number of moles NaCl needed to make 2.5 L of this solution. Convert moles to mass using the molar mass of NaCl (Molar mass = 58.44 g/mol)

$$\text{Mass NaCl} = (0.65 \text{ mol NaCl} / 58.44 \text{ g NaCl} / 2.5 \text{ L}) (1 \text{ mol NaCl} / 58.44 \text{ g NaCl}) = 94.965 = 95 \text{ g NaCl}$$

Add 95 g NaCl to enough water to make 2.5 L of aqueous solution.

- b) Use the relationship $M_{\text{conc}} V_{\text{conc}} = M_{\text{dil}} V_{\text{dil}}$ to find the volume of 2.1 M urea needed.

$$M_{\text{conc}} = 2.1 \text{ M urea } V_{\text{conc}} = 7 \text{ M urea} \Rightarrow 0.3 \text{ M urea } V_{\text{dil}} = 15.5 \text{ L}$$

$$V_{\text{conc}} = M_{\text{dil}} V_{\text{dil}} / M_{\text{conc}} = (0.3 \text{ M}) (15.5 \text{ L}) / (2.1 \text{ M}) = 2.21429 = 2 \text{ L}$$

Add 2 L of 2.1 M urea to enough water to make 15.5 L of solution.

Chapter Section Homework Answers

David Halliday, Robert Resnick, Jearl Walker



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