# BALANCING CHEMICAL EQUATIONS WORKSHEET (REMEMBER PRACTICE MAKES PERFECT)

- 1. \_\_\_\_ Zn (s) + \_\_\_\_ AgNO<sub>3</sub> (aq)  $\rightarrow$  \_\_\_\_ Zn(NO<sub>3</sub>)<sub>2</sub> (aq) + \_\_\_\_ Ag (s) 2. \_\_\_ N<sub>2</sub> (g) + \_\_\_\_ H<sub>2</sub> (g)  $\rightarrow$  \_\_\_\_ NH<sub>3</sub> (g) 3. NaCl (aq) + AgC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> (aq)  $\rightarrow$  NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> (aq) + AgCl (s)
- 4.  $Mg(OH)_2 (aq) + H_3PO_4 (aq) \rightarrow H_2O (l) + Mg_3(PO_4)_2 (aq)$
- 5. \_\_\_\_  $HNO_3$  (aq) + \_\_\_\_ Ni (s)  $\rightarrow$  \_\_\_\_  $Ni(NO_3)_2$  (aq) + \_\_\_\_  $H_2$  (g)
- 6. \_\_\_\_\_ Ba(HCO<sub>3</sub>)<sub>2</sub> (s) → \_\_\_\_\_ BaCO<sub>3</sub> (s) + \_\_\_\_ H<sub>2</sub>O (g) + \_\_\_\_ CO<sub>2</sub> (g)
- 7. \_\_\_\_ BaCl<sub>2</sub> (aq) + \_\_\_\_ Na<sub>2</sub>SO<sub>4</sub> (aq)  $\rightarrow$  \_\_\_\_ NaCl (aq) + \_\_\_\_ BaSO<sub>4</sub> (s)
- 8. \_\_\_\_  $Al_2(CO_3)_3(s) \rightarrow$  \_\_\_\_  $Al_2O_3(s) +$  \_\_\_  $CO_2(g)$
- 9. \_\_\_\_ Ca (s) + \_\_\_ H<sub>2</sub>O (l)  $\rightarrow$  \_\_\_ Ca(OH)<sub>2</sub> (aq) + \_\_\_ H<sub>2</sub> (g)
- 10. \_\_\_\_ LiHCO<sub>3</sub> (s)  $\rightarrow$  \_\_\_\_ Li<sub>2</sub>CO<sub>3</sub> (s) + \_\_\_\_ H<sub>2</sub>O (g) + \_\_\_\_ CO<sub>2</sub> (g)
- 11. \_\_\_\_  $N_2(g) + ___ O_2(g) \rightarrow ___ N_2O_3(g)$
- 12. \_\_\_\_ MgBr<sub>2</sub> (aq) + \_\_\_\_ KOH (aq)  $\rightarrow$  \_\_\_\_ KBr (aq) + \_\_\_ Mg(OH)<sub>2</sub> (s)
- 13. \_\_\_\_ Mn (s) + \_\_\_ CuCl (aq) → \_\_\_ Cu (s) + \_\_\_ MnCl<sub>2</sub> (s)
- 14.  $_$  Zn (s) +  $_$  S<sub>8</sub> (s)  $\rightarrow$  ZnS (s)
- 15. \_\_\_\_ NaOH (aq) + \_\_\_\_  $H_2SO_4$  (aq)  $\rightarrow$  \_\_\_\_  $H_2O$  (l) + \_\_\_\_  $Na_2SO_4$  (aq)
- 16. \_\_\_\_ K (s) + \_\_\_ H<sub>2</sub>O (l)  $\rightarrow$  \_\_\_ KOH (aq) + \_\_\_ H<sub>2</sub> (g)
- 17.  $C_5H_{12}(1) + Q_2(g) \rightarrow H_2O(g) + CO_2(g)$
- 18. \_\_\_\_ KOH (aq) + \_\_\_\_ H<sub>2</sub>CO<sub>3</sub> (aq)  $\rightarrow$  \_\_\_\_ H<sub>2</sub>O (l) + \_\_\_ K<sub>2</sub>CO<sub>3</sub> (aq)
- 19. \_\_\_\_  $C_4H_8O_2(I) + ___ O_2(g) \rightarrow ___ H_2O(g) + ___ CO_2(g)$
- 20.  $Al(s) + S_8(s) \rightarrow Al_2S_3(s)$

#### TYPES OF REACTION

What are the 5 different types of reactions and write the chemical equation given in class?

- 4.
- 5.

For the 20 reactions above write which of the 5 types of reaction it is:

# <u>Chapter 7 Worksheet Answers To Balancinh Chemical</u> <u>Equations</u>

**Alan Trevor Neuff** 

## **Chapter 7 Worksheet Answers To Balancinh Chemical Equations:**

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Workbook with 200+ Reactions to Balance Jake Mariska, 2021-05-08 A chemical equation is something you will encounter every day in chemistry It s a written representation using numbers and symbols of the process that occurs during a chemical reaction Master the art of balancing chemical reactions through examples and practice 10 examples are fully solved step by step with explanations to serve as a guide Over 200 chemical equations provide ample practice Exercises start out easy and grow progressively more challenging and involved Answers to every problem are tabulated at the back of the book A chapter of pre balancing exercises helps develop essential counting skills Opening chapter reviews pertinent concepts and ideas

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