

## Bonding Basics - Covalent Bonds

## Answer Key/Teacher Notes

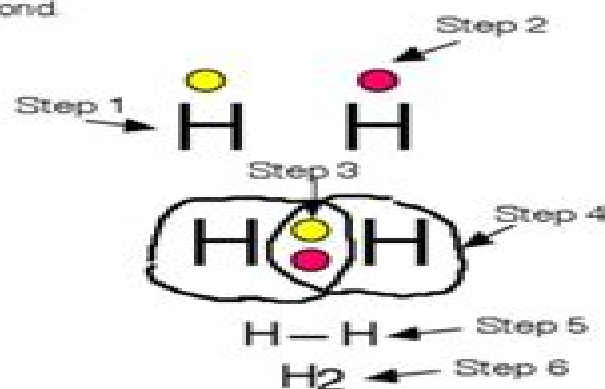
Complete the chart for each element.

Element	# of Protons	# of Electrons	# of Valence Electrons	# of Electrons to Fill Outer Shell
Carbon	6	6	4	4
Chlorine	17	17	7	1
Hydrogen	1	1	1	1
Phosphorus	15	15	5	3
Oxygen	8	8	6	2
Sulfur	16	16	6	2
Nitrogen	7	7	5	3

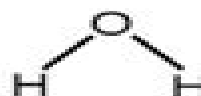
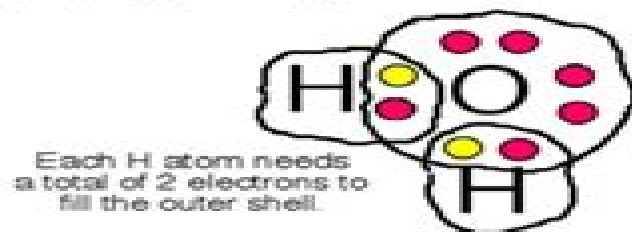
Follow your teacher's directions to complete each covalent bond.

### (1) Hydrogen + Hydrogen (Diatomic Element)

- 1- Write the symbols for each element.
- 2 - Use Fruity Pebbles (or other cereal/candy with more than one color) to create the Lewis structure for each.
- 3 - Rearrange the electrons (or cereal pieces) to pair up electrons from each atom.
- 4 - Draw circles to show the sharing of electrons.
- 5 - Draw the bond structure using symbols and lines. Use one line for each pair of electrons that is shared.
- 6 - Write the chemical formula for each molecule.
- 7 - Have the students use a pencil or crayon to draw the electrons as they remove the pieces of cereal.



### (2) Hydrogen + Oxygen



Students should start with one hydrogen and one oxygen. They should understand that the oxygen atom needs two electrons to fill its outer shell, which means they will need a total of two hydrogen atoms to complete the bond.

T. Trimpe 2002

# Chapter 9 Practice Problems Chemical Bonding Basics

## Answer Key

**Arshad Iqbal**



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