

Chemistry

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(Chapter 3) (Electrochemistry) (Intext Questions)
(Class 12)

Question 3.5:

Calculate the *emf* of the cell in which the following reaction takes place:



Given that $E_{\text{cell}}^{\circ} = 1.05 \text{ V}$

Answer 3.5:

Applying Nernst equation we have:

$$\begin{aligned} E_{\text{(cell)}} &= E_{\text{(cell)}}^{\circ} - \frac{0.0591}{2} \log \frac{[\text{Ni}^{2+}]}{[\text{Ag}^+]^2} \Rightarrow 1.05 - \frac{0.0591}{2} \log \frac{(0.160)}{(0.002)^2} \\ &= 1.05 - 0.02955 \log \frac{0.16}{0.000004} \\ &= 1.05 - 0.02955 \log 4 \times 10^4 \\ &= 1.05 - 0.02955 (\log 10000 + \log 4) \\ &= 1.05 - 0.02955 (4 + 0.6021) \\ &= 0.914 \text{ V} \end{aligned}$$

Question 3.6:

The cell in which the following reactions occurs:



Calculate the standard Gibbs energy and the equilibrium constant of the cell reaction.

Answer 3.6:

Here,

$$n = 2, E_{\text{cell}}^{\circ} = 0.236 \text{ V, } T = 298 \text{ K}$$

We know that:

$$\begin{aligned} \Delta_r G^{\circ} &= -nFE_{\text{cell}}^{\circ} \\ &= -2 \times 96487 \times 0.236 \\ &= -45.54 \text{ kJ mol}^{-1} \\ \Delta_r G^{\circ} &= -2.303RT \log K_c \end{aligned}$$

Again,

$$\begin{aligned} \Rightarrow \log K_c &= -\frac{\Delta_r G^{\circ}}{2.303RT} = -\frac{-45.54 \times 10^3}{2.303 \times 8.314 \times 298} = 7.981 \\ \therefore K_c &= \text{Antilog (7.981)} \\ &= 9.57 \times 10^7 \end{aligned}$$

Question 3.7:

Why does the conductivity of a solution decrease with dilution?

Answer 3.7:

The conductivity of a solution is the conductance of ions present in a unit volume of the solution. The number of ions (responsible for carrying current) decreases when the solution is diluted. As a result, the conductivity of a solution decreases with dilution.

Question 3.8:

Suggest a way to determine the Λ_m° value of water.

Answer 3.8:

Applying Kohlrausch's law of independent migration of ions, the Λ_m° value of water can be determined as follows:

$$\begin{aligned} \Lambda_m^{\circ}(\text{H}_2\text{O}) &= \lambda_{\text{H}^+}^{\circ} + \lambda_{\text{OH}^-}^{\circ} \Rightarrow (\lambda_{\text{H}^+}^{\circ} + \lambda_{\text{Cl}^-}^{\circ}) + (\lambda_{\text{Na}^+}^{\circ} + \lambda_{\text{OH}^-}^{\circ}) - (\lambda_{\text{Na}^+}^{\circ} + \lambda_{\text{Cl}^-}^{\circ}) \\ &\Rightarrow \Lambda_m^{\circ}(\text{HCl}) + \Lambda_m^{\circ}(\text{NaOH}) - \Lambda_m^{\circ}(\text{NaCl}) \end{aligned}$$

Hence, by knowing the Λ_m° values of HCl, NaOH, and NaCl, the Λ_m° value of water can be determined.

Chemistry Chapter Section 3 Note Sheet Answers

Lingsheng Yao



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