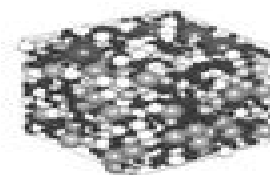




Key

## Density Practice Problems



The density of a substance is a measure of how much mass is packed into a certain volume of the substance. Substances with a high density, like steel, have molecules that are packed together tightly. Substances with a low density, like cork, have fewer molecules packed into the same amount of space.

The density of a substance can be found by dividing its mass by its volume. As long as a substance is homogeneous, the size or shape of the sample doesn't matter. The density will always be the same. This means that a steel paper clip has the same density as a steel girder used to build a bridge.

$$\text{Density} = \text{Mass} \div \text{Volume}$$

$$D = \frac{m}{V}$$

Use the density formula to solve the following problems. Show all work and the answer must have the correct units. Remember that volume can have different forms. A block of ice with a volume of 3 cm<sup>3</sup> would be 3 mL of liquid after being melted.

1. What is the density of CO gas if 0.196 g occupies a volume of 100 ml?

$$m = 0.196 \text{ g}$$

$$V = 100 \text{ ml}$$

$$D = \frac{0.196 \text{ g}}{100 \text{ ml}}$$

$$\text{Answer: } 0.00196 \text{ g/ml} = 0.002 \text{ g/ml}$$

2. A block of wood 3 cm on each side has a mass of 27 g. What is the density of the block? (Hint, don't forget to find the volume of the wood.)

$$V = L \times W \times H$$

$$V = 3 \times 3 \times 3$$

$$V = 27 \text{ cm}^3$$

$$D = \frac{27 \text{ g}}{27 \text{ cm}^3}$$

$$D = 1 \text{ g/cm}^3$$

$$m = 27 \text{ g}$$

$$\text{Answer: } 1 \text{ g/cm}^3$$

3. A student measures a rectangular block of an unknown substance. The block measures 12.1 cm by 17.5 cm by 5.8 cm. The student weighs the block and determines its mass to be 4250.112 g. What is the density of this material? Watch the sig figs.

$$V = L \times W \times H$$

$$V = 12.1 \times 17.5 \times 5.8$$

$$V = 1228.15 \text{ cm}^3$$

$$D = \frac{4250.112 \text{ g}}{1228.15 \text{ cm}^3}$$

$$\text{Answer: } 3.46 \text{ g/cm}^3 = 3.5 \text{ g/cm}^3$$

4. An irregularly shaped stone was lowered into a graduated cylinder holding a volume of water equal to 3.5 ml. The height of the water rose to 8.0 ml. If the mass of the stone was 27 g, what was its density?

$$V = 8.0 \text{ mL} - 3.5 \text{ mL}$$

$$V = 4.5 \text{ mL}$$

$$m = 27 \text{ g}$$

$$D = \frac{27 \text{ g}}{4.5 \text{ mL}} = 6$$

$$\text{Answer: } 6.0 \text{ g/mL}$$

# Density Practice Problem Answer Key Chapter 7

**Esther Muscari Lin, Oncology Nursing  
Society**



## **Density Practice Problem Answer Key Chapter 7:**

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