

# Chapter 7 Worksheet #1

## Balancing Chemical Equations

Balance the equations below:

- 1)  $\underline{\quad\quad} \text{N}_2 + \underline{3} \text{H}_2 \rightarrow \underline{2} \text{NH}_3$
- 2)  $\underline{2} \text{KClO}_3 \rightarrow \underline{2} \text{KCl} + \underline{3} \text{O}_2$
- 3)  $\underline{2} \text{NaCl} + \underline{\quad\quad} \text{F}_2 \rightarrow \underline{2} \text{NaF} + \underline{\quad\quad} \text{Cl}_2$
- 4)  $\underline{2} \text{H}_2 + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{2} \text{H}_2\text{O}$
- 5)  $\underline{\quad\quad} \text{Pb}(\text{OH})_2 + \underline{2} \text{HCl} \rightarrow \underline{2} \text{H}_2\text{O} + \underline{\quad\quad} \text{PbCl}_2$
- 6)  $\underline{2} \text{AlBr}_3 + \underline{3} \text{K}_2\text{SO}_4 \rightarrow \underline{6} \text{KBr} + \underline{\quad\quad} \text{Al}_2(\text{SO}_4)_3$
- 7)  $\underline{\quad\quad} \text{CH}_4 + \underline{2} \text{O}_2 \rightarrow \underline{\quad\quad} \text{CO}_2 + \underline{2} \text{H}_2\text{O}$
- 8)  $\underline{\quad\quad} \text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- 9)  $\underline{\quad\quad} \text{C}_6\text{H}_{18} + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{8} \text{CO}_2 + \underline{9} \text{H}_2\text{O} \quad \checkmark$
- 10)  $\underline{\quad\quad} \text{FeCl}_3 + \underline{3} \text{NaOH} \rightarrow \underline{\quad\quad} \text{Fe}(\text{OH})_3 + \underline{3} \text{NaCl}$
- 11)  $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- 12)  $\underline{\quad\quad} \text{Na} + \underline{\quad\quad} \text{H}_2\text{O} \rightarrow \underline{\quad\quad} \text{NaOH} + \underline{\quad\quad} \text{H}_2 \quad \checkmark$
- 13)  $\underline{2} \text{Ag}_2\text{O} \rightarrow \underline{4} \text{Ag} + \underline{\quad\quad} \text{O}_2$
- 14)  $\underline{\quad\quad} \text{S}_8 + \underline{12} \text{O}_2 \rightarrow \underline{8} \text{SO}_3$
- 15)  $\underline{6} \text{CO}_2 + \underline{6} \text{H}_2\text{O} \rightarrow \underline{\quad\quad} \text{C}_6\text{H}_{12}\text{O}_6 + \underline{9} \text{O}_2$
- 16)  $\underline{2} \text{K} + \underline{\quad\quad} \text{MgBr}_2 \rightarrow \underline{2} \text{KBr} + \underline{\quad\quad} \text{Mg}$
- 17)  $\underline{2} \text{HCl} + \underline{\quad\quad} \text{CaCO}_3 \rightarrow \underline{\quad\quad} \text{CaCl}_2 + \underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{CO}_2$
- 18)  $\underline{\quad\quad} \text{HNO}_3 + \underline{\quad\quad} \text{NaHCO}_3 \rightarrow \underline{\quad\quad} \text{NaNO}_3 + \underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{CO}_2$
- 19)  $\underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{\quad\quad} \text{H}_2\text{O}_2$
- 20)  $\underline{\quad\quad} \text{NaBr} + \underline{\quad\quad} \text{CaF}_2 \rightarrow \underline{\quad\quad} \text{NaF} + \underline{\quad\quad} \text{CaBr}_2$
- 21)  $\underline{\quad\quad} \text{H}_2\text{SO}_4 + \underline{\quad\quad} \text{NaNO}_2 \rightarrow \underline{\quad\quad} \text{HNO}_2 + \underline{\quad\quad} \text{Na}_2\text{SO}_4$

# Chapter 7 Worksheet 1 Balancing Chemical Equations

## Answer Key

**Holt Rinehart & Winston**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

## Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key:

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### **Table of Contents Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key**

1. Understanding the eBook Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - The Rise of Digital Reading Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - User-Friendly Interface

## **Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key**

---

4. Exploring eBook Recommendations from Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Personalized Recommendations
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key User Reviews and Ratings
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key and Bestseller Lists
5. Accessing Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key Free and Paid eBooks
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key Public Domain eBooks
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key eBook Subscription Services
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key Budget-Friendly Options
6. Navigating Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key Compatibility with Devices
  - Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Highlighting and Note-Taking Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Interactive Elements Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
8. Staying Engaged with Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
9. Balancing eBooks and Physical Books Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Setting Reading Goals Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Fact-Checking eBook Content of Chapter 7 Worksheet 1 Balancing Chemical Equations Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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