

Chapter 7 Worksheet #1

Balancing Chemical Equations

Balance the equations below:

- 1) $\underline{\quad\quad} \text{N}_2 + \underline{3} \text{H}_2 \rightarrow \underline{2} \text{NH}_3$
- 2) $\underline{2} \text{KClO}_3 \rightarrow \underline{2} \text{KCl} + \underline{3} \text{O}_2$
- 3) $\underline{2} \text{NaCl} + \underline{\quad\quad} \text{F}_2 \rightarrow \underline{2} \text{NaF} + \underline{\quad\quad} \text{Cl}_2$
- 4) $\underline{2} \text{H}_2 + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{2} \text{H}_2\text{O}$
- 5) $\underline{\quad\quad} \text{Pb}(\text{OH})_2 + \underline{2} \text{HCl} \rightarrow \underline{2} \text{H}_2\text{O} + \underline{\quad\quad} \text{PbCl}_2$
- 6) $\underline{2} \text{AlBr}_3 + \underline{3} \text{K}_2\text{SO}_4 \rightarrow \underline{6} \text{KBr} + \underline{\quad\quad} \text{Al}_2(\text{SO}_4)_3$
- 7) $\underline{\quad\quad} \text{CH}_4 + \underline{2} \text{O}_2 \rightarrow \underline{\quad\quad} \text{CO}_2 + \underline{2} \text{H}_2\text{O}$
- 8) $\underline{\quad\quad} \text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- 9) $\underline{\quad\quad} \text{C}_6\text{H}_{14} + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{8} \text{CO}_2 + \underline{9} \text{H}_2\text{O} \quad \checkmark$
- 10) $\underline{\quad\quad} \text{FeCl}_3 + \underline{3} \text{NaOH} \rightarrow \underline{\quad\quad} \text{Fe}(\text{OH})_3 + \underline{3} \text{NaCl}$
- 11) $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- 12) $\underline{\quad\quad} \text{Na} + \underline{\quad\quad} \text{H}_2\text{O} \rightarrow \underline{\quad\quad} \text{NaOH} + \underline{\quad\quad} \text{H}_2 \quad \checkmark$
- 13) $\underline{2} \text{Ag}_2\text{O} \rightarrow \underline{4} \text{Ag} + \underline{\quad\quad} \text{O}_2$
- 14) $\underline{\quad\quad} \text{S}_8 + \underline{12} \text{O}_2 \rightarrow \underline{8} \text{SO}_3$
- 15) $\underline{6} \text{CO}_2 + \underline{6} \text{H}_2\text{O} \rightarrow \underline{\quad\quad} \text{C}_6\text{H}_{12}\text{O}_6 + \underline{9} \text{O}_2$
- 16) $\underline{2} \text{K} + \underline{\quad\quad} \text{MgBr}_2 \rightarrow \underline{2} \text{KBr} + \underline{\quad\quad} \text{Mg}$
- 17) $\underline{2} \text{HCl} + \underline{\quad\quad} \text{CaCO}_3 \rightarrow \underline{\quad\quad} \text{CaCl}_2 + \underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{CO}_2$
- 18) $\underline{\quad\quad} \text{HNO}_3 + \underline{\quad\quad} \text{NaHCO}_3 \rightarrow \underline{\quad\quad} \text{NaNO}_3 + \underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{CO}_2$
- 19) $\underline{\quad\quad} \text{H}_2\text{O} + \underline{\quad\quad} \text{O}_2 \rightarrow \underline{\quad\quad} \text{H}_2\text{O}_2$
- 20) $\underline{\quad\quad} \text{NaBr} + \underline{\quad\quad} \text{CaF}_2 \rightarrow \underline{\quad\quad} \text{NaF} + \underline{\quad\quad} \text{CaBr}_2$
- 21) $\underline{\quad\quad} \text{H}_2\text{SO}_4 + \underline{\quad\quad} \text{NaNO}_2 \rightarrow \underline{\quad\quad} \text{HNO}_2 + \underline{\quad\quad} \text{Na}_2\text{SO}_4$

Chapter 7 Worksheet 1 Balancing Chemical Equations Answers

Christian Drosten



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