Oxoacids (O—H bond): contain H, O, and a central non-metal Acidity determined by the polarity of the O—H bond:

 central atom from the same group of the periodic table, same oxidation number

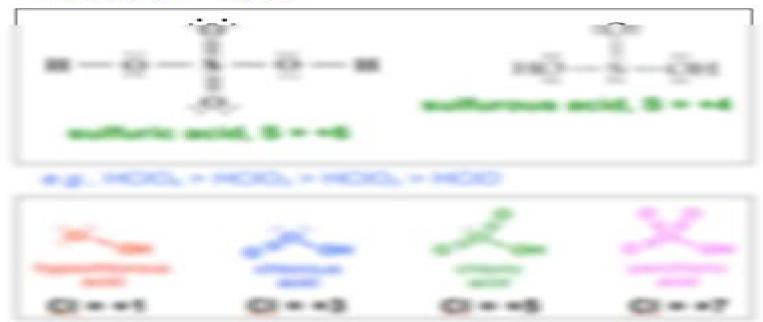
Relative strength increases with increased electronegativity of the central atom.

HClO<sub>3</sub> > HBrO<sub>3</sub>

H<sub>2</sub>SO<sub>4</sub> > H<sub>2</sub>SeO<sub>4</sub>

2. same central atom, different oxidation number (# of O's)

Higher oxidation number = stronger acid.



# **Chapter 8 3 Solution Acids And Bases Abswer Key**

**Disha Experts** 

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