

## Chemistry--Unit 3: Chemical Quantities

### Practice Problems

#### I. The Mole: A Measurement of Matter

- 1) What is the gram molecular mass of sucrose ( $C_{12}H_{22}O_{11}$ )? (342.34 g/mol)
- 2) Calculate the gram formula mass of  $KMnO_4$  and  $Ca_3(PO_4)_2$ .  
( $KMnO_4 = 158.04$  g/mol,  $Ca_3(PO_4)_2 = 310.18$  g/mol)
- 3) How many moles is  $3.52 \times 10^{24}$  molecules of water?

$$3.52 \times 10^{24} \text{ molecules} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}} = 5.85 \text{ mol}$$

- 4) How many atoms of zinc are in 0.60 mol of zinc?

$$0.60 \text{ mol Zn} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} = 3.61 \times 10^{23} \text{ atoms}$$

#### II. Mole-Mass and Mole-Volume Relationships

- 5) Calculate the mass in grams of each of the following.

- a. 8.0 mol lead(II) oxide, ( $PbO$ )

$$8.0 \text{ mol PbO} \times \frac{223.2 \text{ g}}{1 \text{ mol}} = 1785.6 \text{ g}$$

- b. 0.75 mol hydrogen sulfide ( $H_2S$ )

$$0.75 \text{ mol } H_2S \times \frac{34.08 \text{ g}}{1 \text{ mol}} = 25.56 \text{ g}$$

- c.  $1.50 \times 10^{25}$  molecules of molecular oxygen ( $O_2$ )

$$1.50 \times 10^{25} \text{ molecules} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}} \times \frac{32.00 \text{ g}}{1 \text{ mole}} = 797.3 \text{ g}$$

- d.  $2.30 \times 10^{30}$  atoms of iridium

$$2.30 \times 10^{30} \text{ atoms} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}} \times \frac{192.22 \text{ g}}{1 \text{ mole}} = 7.34 \times 10^5 \text{ g}$$

- 6) Calculate the number of moles in each of the following.

- a. 0.50 g sodium bromide ( $NaBr$ )

$$0.50 \text{ g} \times \frac{1 \text{ mole}}{102.89 \text{ g}} = 0.0049 \text{ mol}$$

# Chapter Chemical Quantities Worksheet Answers

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## **Table of Contents Chapter Chemical Quantities Worksheet Answers**

1. Understanding the eBook Chapter Chemical Quantities Worksheet Answers
  - The Rise of Digital Reading Chapter Chemical Quantities Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Chemical Quantities Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter Chemical Quantities Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Chemical Quantities Worksheet Answers
  - Personalized Recommendations
  - Chapter Chemical Quantities Worksheet Answers User Reviews and Ratings
  - Chapter Chemical Quantities Worksheet Answers and Bestseller Lists
5. Accessing Chapter Chemical Quantities Worksheet Answers Free and Paid eBooks
  - Chapter Chemical Quantities Worksheet Answers Public Domain eBooks

- Chapter Chemical Quantities Worksheet Answers eBook Subscription Services
- Chapter Chemical Quantities Worksheet Answers Budget-Friendly Options
- 6. Navigating Chapter Chemical Quantities Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter Chemical Quantities Worksheet Answers Compatibility with Devices
  - Chapter Chemical Quantities Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter Chemical Quantities Worksheet Answers
  - Highlighting and Note-Taking Chapter Chemical Quantities Worksheet Answers
  - Interactive Elements Chapter Chemical Quantities Worksheet Answers
- 8. Staying Engaged with Chapter Chemical Quantities Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter Chemical Quantities Worksheet Answers
- 9. Balancing eBooks and Physical Books Chapter Chemical Quantities Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter Chemical Quantities Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Chemical Quantities Worksheet Answers
  - Setting Reading Goals Chapter Chemical Quantities Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Chemical Quantities Worksheet Answers
  - Fact-Checking eBook Content of Chapter Chemical Quantities Worksheet Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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