

Section 6.1: Organizing the Elements

- Group 1 on the periodic table is called **Alkali Metals**.
- Group 2 on the periodic table is called **Alkaline Earth Metals**.
- Elements in groups 3-12 are called **Transition Metals**.
- Group 17 on the periodic table is called **Halogens**.
- Group 18 on the periodic table is called **Noble Gases**.
- Each period on the periodic table corresponds to a **principal energy level**.
- The 2 series that make up the inner transition metals are **Lanthanides** and **Actinides**.
- Complete the following chart:

Element	Symbol	Group #	Period #	Group Name	Metal, Metalloid, Nonmetal
Zirconium	Zr	4	5	Transition Metals	Metal
Radium	Ra	2	7	Alkaline Earth Metals	Metal
Bismuth	Bi	15	6	Nitrogen Group	Metal
Gallium	Ga	13	4	Boron Group	Metal

- What are triads? **Sets of 3 elements with similar properties**. Whose idea were they? **Dobereiner**.
- The modern periodic table is arranged by increasing **atomic number**. Whose idea was this? **Moseley**.
- Mendeleev** developed the first periodic table, which was arranged by increasing **atomic mass**.
- Determine which of the following pairs will have similar properties (Yes or no):
 - Ca and Be? **Yes – same group (Group 2 – Alkaline Earth Metals)**
 - Sr and Y? **No – not the same group**
 - C and Ge? **Yes – same group (Group 14 – Carbon Group)**

Section 6.3: Periodic Trends

- As you go across a period, the atomic radius (size) **decreases** because **the attraction between protons and electrons increases, which makes the atom smaller (opposite of tug of war)**.
- As you go down a group, the atomic radius (size) **increases** because **the number of principal energy levels increases**.
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:

a. Li	Ne	F	B
b. N	Bi	S	P
c. Na	Sr	Al	Cl
- Ionization energy is the energy required to **remove** an electron.

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