Chapter 2 Study Guide

CHAPTER 2 The Chemical Basis of Life 1: Atoms, Molecules, and Water

- Describe the location, charge and mass of protons, neutrons and electrons, within an atom.
 - Neutrons and Protons are densely packed in the Nucleus.
 - b. Electrons move about the Nucleus in Orbitals.
 - c. Neutrons and Protons have a mass of one.
- Be able to calculate the Atomic Weight of an atom/element, if the number of protons, electrons and neutrons is known.
 - a. Atomic Mass Number = Protons + Neutrons
- Be able to calculate the Atomic Number of an atom/element, if the number of protons, electrons and neutrons is known.
 - a. Atomic Number Number of Protons
- Be able to write the symbols, atomic numbers, and mass numbers of the biologically important elements (C, H, O, & N).
 - a. C → Carbon, 12 Protons, 12 011 Mass
 - b. H → Hydrogen, 1 Proton, 1.008 Mass
 - c. O → Curygen, 8 Protons, 15.999 Mass
 - d. N → Nitropen, 7 Protons, 14 007 Mass
- Be able to explain how isotopes of an element differ in terms of the number of electrons, neutrons and protons.
 - a. Isotopes differ in the number of neutrons in the Nucleus.
 - The number of Protons does not change.
 - c. The number of Electrons does not change.
- Be able to correctly place any number of elections in an atom's first two energy level shells.
 - a. First Energy Level requires 2 Electrons.
 - Any Energy Levels after the first require 8 Electrons
- Differentiate between atoms, molecules and ions.
 - a. Atoms are single neutral particles.
 - b. Molecules are neutral particles made of two or more atoms bonded together.
 - c. Ions are positively or negatively charged particles (atoms)
- Differentiate between molecules and compounds.
 - All compounds are molecules, but not all molecules are compounds.
 - b. Molecules are made up of any two or more atoms bonded together, but Compounds are made up of two or more different atoms bonded together.
- What is a valence electron?
 - a. Valence Electrons are Electrons located in the outermost Energy Shell of an atom. Since they are the farthest from the Nucleus, they can participate in bonds and reactions, therefore Valence Electrons determine an atom's reactivity, electronegativity, and the number of bonds it can form.
- Describe a covalent chemical bond, and what holds the two atoms together.
 - a. A Covalent Bond is formed through the equal sharing of valence Electrons in two atoms, with the sharing process enabling each atom to fill its outer valence electrons.
- Differentiate between a polar and non-polar covalent bond.
 - a. A Polar Covalent Band results if the Electronegativity difference from a band is between 0.5-1.6, meaning one end of the band is partly negative while the other is partly positive.
 - b. A Non-Polar Covalent Bond results if the Electronegativity difference from a bond is between 0.2-0.5.
- 12. What is a Hydrogen Bond, and what holds the two atoms together?
 - a. A Hydrogen B and occurs only with Polar Molecules, as the partial opposite charges of the molecules will attract one another, usually placing a Hydrogen atom as the bridge between the molecules.
- What is an ionic bond, and what holds the two atoms together.
 - a. An lonic Bond is a bond that is formed from the transfer of Electrons from one atom to another, therefore making one element have more Electrons that Protons and the other have fewer Electrons than Protons.
 - Essentially, the more Electronegative atom strips Electrons from another atom to complete its outer-most valence shell.
- 14. What is a van der Waals interaction (bond) and what holds the two atoms together?

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Ying-Ying Zheng

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Table of Contents Ch 2study Guide Transmutation Answers

- 1. Understanding the eBook Ch 2study Guide Transmutation Answers
 - The Rise of Digital Reading Ch 2study Guide Transmutation Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ch 2study Guide Transmutation Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 2study Guide Transmutation Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ch 2study Guide Transmutation Answers
 - Personalized Recommendations
 - o Ch 2study Guide Transmutation Answers User Reviews and Ratings
 - Ch 2study Guide Transmutation Answers and Bestseller Lists
- 5. Accessing Ch 2study Guide Transmutation Answers Free and Paid eBooks
 - Ch 2study Guide Transmutation Answers Public Domain eBooks
 - Ch 2study Guide Transmutation Answers eBook Subscription Services
 - Ch 2study Guide Transmutation Answers Budget-Friendly Options
- 6. Navigating Ch 2study Guide Transmutation Answers eBook Formats

- o ePub, PDF, MOBI, and More
- Ch 2study Guide Transmutation Answers Compatibility with Devices
- Ch 2study Guide Transmutation Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Ch 2study Guide Transmutation Answers
 - Highlighting and Note-Taking Ch 2study Guide Transmutation Answers
 - Interactive Elements Ch 2study Guide Transmutation Answers
- 8. Staying Engaged with Ch 2study Guide Transmutation Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 2study Guide Transmutation Answers
- 9. Balancing eBooks and Physical Books Ch 2study Guide Transmutation Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 2study Guide Transmutation Answers
- 10. Overcoming Reading Challenges
 - $\circ\,$ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 2study Guide Transmutation Answers
 - Setting Reading Goals Ch 2study Guide Transmutation Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 2study Guide Transmutation Answers
 - Fact-Checking eBook Content of Ch 2study Guide Transmutation Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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equitable a compromise that is fair to both factions 7 being in accordance with relative merit or significance she wanted to