

Chapter 4 Atomic Structure

Section 4.1 Studying Atoms**(pages 100-105)***This section discusses the development of atomic models.***Reading Strategy (page 100)**

Summarizing As you read, complete the table about atomic models. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Atomic Models		
Scientist	Evidence	Model
dalton	Ratio of masses in compounds	Indivisible, solid spheres
thomson	Deflected beam	Negative charges evenly scattered through positively charged mass of matter
Rutherford	Deflection of alpha particles passing through gold foil	Positive, dense nucleus

Ancient Greek Models of Atoms (page 100)

- Democritus named the smallest particles of matter atoms because they could not be divided.
- List the four elements that Aristotle included in his model of matter.
 - earth
 - fire
 - water
 - air

Dalton's Atomic Theory (page 101)

- Is the following sentence true or false? John Dalton gathered evidence for the existence of atoms by measuring the masses of elements that reacted to form compounds. true
- What theory did Dalton propose to explain why the elements in a compound always join in the same way? All elements are composed of atoms. In a particular compound, atoms of different elements always combine the same way. Compounds contain atoms of more than one element.
- Circle the letters of the sentences that represent the main points of Dalton's theory of atoms.
 - All elements are composed of atoms.
 - In a particular compound, atoms of different elements always combine the same way.
 - All atoms have the same mass.
 - Compounds contain atoms of more than one element.

Ch 4 Guided Atomic Structure Answers

Karen C. Timberlake



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