

- \_\_\_\_\_ 13. If the spin of one electron in an orbital is clockwise, what is the spin of the other electron in that orbital?
- a. zero  
b. clockwise  
c. counterclockwise  
d. both clockwise and counterclockwise
- \_\_\_\_\_ 14. What types of atomic orbitals are in the third principal energy level?
- a. *s* and *p* only  
b. *p* and *d* only  
c. *s*, *p*, and *d* only  
d. *s*, *p*, *d*, and *f*
- \_\_\_\_\_ 15. What is the next atomic orbital in the series 1*s*, 2*s*, 2*p*, 3*s*, 3*p*?
- a. 2*d*  
b. 3*d*  
c. 3*f*  
d. 4*s*
- \_\_\_\_\_ 16. What is the number of electrons in the outermost energy level of an oxygen atom?
- a. 2  
b. 4  
c. 6  
d. 8
- \_\_\_\_\_ 17. What is the electron configuration of potassium?
- a.  $1s^2 2s^2 2p^2 3s^2 3p^2 4s^1$   
b.  $1s^2 2s^2 2p^{10} 3s^2 3p^3$   
c.  $1s^2 2s^2 3s^2 3p^6 3d^6$   
d.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- \_\_\_\_\_ 18. How many unpaired electrons are in a sulfur atom (atomic number 16)?
- a. 0  
b. 1  
c. 2  
d. 3
- \_\_\_\_\_ 19. Which is the correct electron dot diagram for the element Silicon (Si)?
- A                  B                  C                  D
- $\cdot \ddot{\text{Si}} \cdot$        $:\ddot{\text{Si}}:$        $\dot{\text{Si}}\cdot$        $: \text{Si} :$
- a. \_\_\_\_\_  
b. \_\_\_\_\_  
c. \_\_\_\_\_  
d. \_\_\_\_\_
- \_\_\_\_\_ 20. How many valence electrons are present in silicon?
- a. Two  
b. Four  
c. Six  
d. Eight

**Completion:** Complete each statement.

21. The \_\_\_\_\_ principle states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins.
22. The arrangement of electrons in an atom is called the \_\_\_\_\_ of an atom.
23. The element \_\_\_\_\_ has the following electron configuration:  $[\text{Rn}]5f^6 6d^1 7s^2$ .
24. The *f*-block elements are also referred to as the lanthanides and \_\_\_\_\_.
25. Xenon has \_\_\_\_\_ valence electrons.
26. Aluminum has \_\_\_\_\_ valence electrons.

# Ch 5 Test Electrons In Atoms Answers

**JG Myers**



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### **Table of Contents Ch 5 Test Electrons In Atoms Answers**

1. Understanding the eBook Ch 5 Test Electrons In Atoms Answers
  - The Rise of Digital Reading Ch 5 Test Electrons In Atoms Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ch 5 Test Electrons In Atoms Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ch 5 Test Electrons In Atoms Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 5 Test Electrons In Atoms Answers
  - Personalized Recommendations
  - Ch 5 Test Electrons In Atoms Answers User Reviews and Ratings
  - Ch 5 Test Electrons In Atoms Answers and Bestseller Lists

5. Accessing Ch 5 Test Electrons In Atoms Answers Free and Paid eBooks
  - Ch 5 Test Electrons In Atoms Answers Public Domain eBooks
  - Ch 5 Test Electrons In Atoms Answers eBook Subscription Services
  - Ch 5 Test Electrons In Atoms Answers Budget-Friendly Options
6. Navigating Ch 5 Test Electrons In Atoms Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ch 5 Test Electrons In Atoms Answers Compatibility with Devices
  - Ch 5 Test Electrons In Atoms Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ch 5 Test Electrons In Atoms Answers
  - Highlighting and Note-Taking Ch 5 Test Electrons In Atoms Answers
  - Interactive Elements Ch 5 Test Electrons In Atoms Answers
8. Staying Engaged with Ch 5 Test Electrons In Atoms Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ch 5 Test Electrons In Atoms Answers
9. Balancing eBooks and Physical Books Ch 5 Test Electrons In Atoms Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ch 5 Test Electrons In Atoms Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ch 5 Test Electrons In Atoms Answers
  - Setting Reading Goals Ch 5 Test Electrons In Atoms Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ch 5 Test Electrons In Atoms Answers
  - Fact-Checking eBook Content of Ch 5 Test Electrons In Atoms Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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