CLASSIFYING CHEMICAL REACTIONS



Classify the following reactions as synthesis, decomposition, single replacement or double replacement. De composition 2KCIO, → 2KCI + 3O, Double Replacement HCI + NaOH → NaCI + H_{*}O Single Replacement. Mg + 2HCl → MgCl, + H, Synthesis 2H₂ + O₂ → 2H₂O Single Replacement 2AI + 3NIBr₂ → 2AIBr₃ + 3NI Synthesis 4AI + 3O₂ → 2AI₂O₃ Decomposition 2NaCl → 2Na + Cl_a Single Replacement CaCl₂ + F₂ → CaF₂ + Cl₂ Double Replacement AgNO₃ + KCI → AgCI + KNO₃ Synthesis N₂ + 3H₂ → 2NH₃ Decomposition 2H₂O₃ → 2H₂O + O₃ 12. (NH,)250, + Ba(NO3)2 - Baso, + 2NH,NO, Double Replacement Single Replacement 13. Mgl₂ + Br₂ → MgBr₂ + l₂ Synthesis 14. SO, + H₂O → H₂SO₄

7 11 0 1

<u>Classifying Chemical Equations Classify As Synthesis</u> <u>Answers</u>

Ralph A. Burns

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