

Activity 10

Classifying reactions

On the Go: 10 min

Types of chemical reactions and their characteristics



Introduction: Chemical reactions show how **compounds** and **elements** react with one another. An element is an substance consisting of only one of a kind. Such as the elements of hydrogen, oxygen and fluorine. A compound is a substance made of more than one kind of atoms, such as water, which is made up of hydrogen.

Question: How are chemical reactions classified?

1. Study these chemical reactions and be classified as one of four types. Using the chemical equations in the boxes on the right, match the following definitions to the type of reaction.

☒ **A** One reaction in which there is one reactant and one product.

a. Synthesis

☒ **A** A reaction in which one reactant breaks down into two or more products.

a. Decomposition

☒ **A** Two or more reactants combine to form one product.

c. Single replacement

☒ **A** Two compounds react to form two different compounds.

b. Double replacement

☒ **A** A compound reacts with an element to form a new compound and a different element.

d. Combustion

2. Study these reactions and classify the chemical reactions below. Some reactions may already be classified in the space on the right. Identify the reaction as a synthesis, decomposition, single replacement, or double replacement reaction.

a. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

Double

b. $\text{NaCl} \rightarrow \text{Na} + \text{Cl}_2$

Single

c. ☒ $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

Synthesis

d. ☒ $\text{NaOH} + \text{NaOH} \rightarrow \text{NaOH} + \text{NaOH}$

Synthesis

e. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2$

Combustion

Balancing Equations Gizmo Key

S Ashworth

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