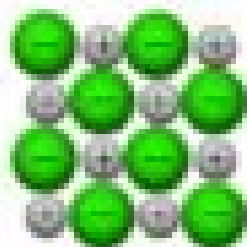


<https://www.youtube.com/watch?v=1j0m0p0p0p0>



<https://www.youtube.com/watch?v=1j0m0p0p0p0>

The Chemistry

Ionic, Covalent, and metallic compounds each have physical properties that relate to their structures, so there are some tests we can do to help identify the type of bonding in a substance.

Melting Points

An ionic bond is made when ions of opposite charges come together to create a **crystal** lattice structure. In order to melt an ionic compound, the temperature has to rise until the ions move with enough force to break away from the oppositely charged ions around them.

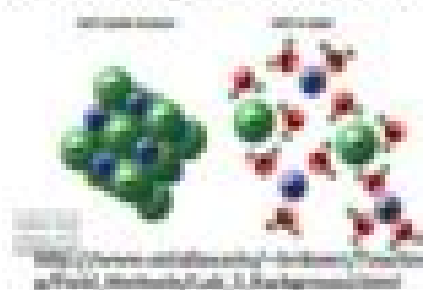
Covalent Compounds don't break apart when they melt, so they only have to reach a temperature strong enough to pull them away from the other molecules to melt them. This is why most small covalent molecules are liquid or gas at room temperature.



<https://www.youtube.com/watch?v=1j0m0p0p0p0>

11/11/2020

Solubility & Conductivity



Solubility and Conductivity

Ionic compounds **dissolve very easily** in water to create an aqueous solution. Since water is a polar molecule, it attracts the ions and surrounds each one as the compound dissociates in the water. Once water is full of charged particles, it can conduct electricity very easily. Tap water conducts electricity because the chemicals used to treat the water are ions, so it makes the water into a conductor.



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Ying Liu



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