

CHEMISTRY 1 – SEMESTER 2

REVIEW GUIDE

Chemical Reactions

Chemical Reaction: process by which one or more substances are changed into different substances.

Total mass of reactants must equal total mass of products

Evidence of reactions:

- Heat
- Light
- Gas
- Precipitate
- Color change

Reactants: substances being reacted

Products: substances produced by reaction

Chemical Equation: a representation of a chemical reaction using symbols and formulas

$\text{CH}_4 (\text{g}) + 2 \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2 \text{H}_2\text{O} (\text{l})$ Reaction between methane and oxygen to produce carbon dioxide and water

Coefficients: relative number of moles

1 mole of CH_4 + 2 moles of O_2 → 1 mole CO_2 + 2 moles H_2O

$3 \text{Fe} (\text{s}) + 4 \text{H}_2\text{O} (\text{g}) \rightarrow \text{Fe}_3\text{O}_4 (\text{s}) + 4 \text{H}_2 (\text{g})$

States of matter: solid/liquid/gas

Solid iron + gas water → solid iron oxide + gas hydrogen

Word Equations: chemical reactions represented in words without formulas or symbols (or coefficients)

Sodium oxide + water → sodium hydroxide

$\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{NaOH}$ Unbalanced

Balanced Equations:

- Correct formulas
- Correct number of moles

Never change subscripts

Polyatomic ions are single units

Count all atoms at the end to make sure

Unbalanced: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Unbalanced: $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

Balanced: $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$

Balanced: $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$

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Susan McMurry



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Table of Contents Covalent Review Guide

1. Understanding the eBook Covalent Review Guide
 - The Rise of Digital Reading Covalent Review Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Review Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Review Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Review Guide
 - Personalized Recommendations

- Covalent Review Guide User Reviews and Ratings
- Covalent Review Guide and Bestseller Lists
- 5. Accessing Covalent Review Guide Free and Paid eBooks
 - Covalent Review Guide Public Domain eBooks
 - Covalent Review Guide eBook Subscription Services
 - Covalent Review Guide Budget-Friendly Options
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 - Covalent Review Guide Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Covalent Review Guide
 - Highlighting and Note-Taking Covalent Review Guide
 - Interactive Elements Covalent Review Guide
- 8. Staying Engaged with Covalent Review Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Review Guide
- 9. Balancing eBooks and Physical Books Covalent Review Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Review Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Review Guide
 - Setting Reading Goals Covalent Review Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Review Guide
 - Fact-Checking eBook Content of Covalent Review Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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