

B = Bonding & Molecular Structure

Table B.1: Some Average Single and Multiple Bond Energies (kJ/mol)

	H	Li	Na	K	Rb	Cs	F	Cl	Br	I	At
H-H	436						438	242	232	213	
H-Cl		431	402	375	359	343					
H-Br			366	343	325	312					
H-I				300	281	267					
H-At											
Li-H							404	357	339	324	
Na-H							391	335	316	302	
K-H							380	324	305	291	
Rb-H							369	313	294	280	
Cs-H							358	302	283	269	
Li-F							593	505	484	464	
Na-F							578	489	468	448	
K-F							563	474	453	433	
Rb-F							548	459	438	418	
Cs-F							533	444	423	403	
Li-Cl							403	339	325	310	
Na-Cl							381	317	303	288	
K-Cl							369	305	291	276	
Rb-Cl							358	294	280	265	
Cs-Cl							347	283	269	254	
Li-Br							366	301	287	272	
Na-Br							344	280	266	251	
K-Br							332	268	254	239	
Rb-Br							321	257	243	228	
Cs-Br							310	246	232	217	
Li-I							300	235	221	206	
Na-I							288	223	209	194	
K-I							276	211	197	182	
Rb-I							265	200	186	171	
Cs-I							254	189	175	160	
Li-At							299	234	220	205	
Na-At							287	222	208	193	
K-At							275	210	196	181	
Rb-At							264	199	185	170	
Cs-At							253	188	174	159	

Multiple Bonds			
C-C	346	C=C	614
C=C	614	C≡C	839
C≡C	839	C≡N	891
C≡N	891	C≡O	1072
C≡O	1072	N≡N	945

1. Write the balanced equation for the combustion of 1 mole of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, completely in O_2 .



2. Draw the structural formula showing Lewis structures for each of the molecules involved.

B = 1.07 g/mol C = 12.01 g/mol H = 1.01 g/mol O = 16.00 g/mol N = 14.01 g/mol



3. Calculate the bond energies of the reactants.

Bond	kJ/mol	kJ/mol
C-H	413	2478
C-O	336	3792
O=O	498	1195

- Calculate the bond energies of the products.

Bond	kJ/mol	kJ/mol
C=O	799	3196
O-H	463	2728

Total Bond Energy (Reactants)	-4627 kJ	Total Bond Energy (Products)	-5974 kJ
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4. Calculate the Heat Energy released. Bond Energy values.

What is the Heat Energy released (combustion)? -1347 kJ

Bonding And Molecular Structure Worksheet Answers South Pasadena

Harry B. Gray



Bonding And Molecular Structure Worksheet Answers South Pasadena:

Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1989 Designed for use in inorganic physical and quantum chemistry courses this textbook includes numerous questions and problems at the end of each chapter and an Appendix with answers to most of the problems

Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1965

Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973

Structure and Bonding Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes

Chemical Bonding and the Geometry of Molecules George Eugene Ryschkewitsch, 1965

Bonding and Molecular Structure Russell S. Drago, 1971

The Shape and Structure of Molecules Charles Alfred Coulson, 1973

A Pictorial Approach to Molecular Bonding John G. Verkade, 2012-12-06

With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example

Molecules and the Chemical Bond Henry A. Bent, 2011

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Molecular Structure and Energetics; Volume 1: Chemical Bonding Models, 1986 **The Determination of Molecular Structure** Peter Jaffrey Wheatley, 1962

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