

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	None
2) $\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	None
3) $\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

Double Replacement Reaction And Solubility Lab Key

M Walker

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Whole-class Inquiry Dennis W. Smithenry,Joan Gallager-Bolos,2009 In response to requests from science education professionals this is the perfect vehicle for implementing and assessing this concept of whole class inquiry in your classroom This is a must have package for preservice and inservice middle and high school science teachers Prentice Hall Chemistry ,2000 Chemistry Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 **Improving Student Comprehension in Chemistry Laboratories** Tracy Lynn Haroff,2006

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of synthetic chemists and pharmaceutical and natural product researchers These authors have reviewed the tremendous body of literature in the field in order to compile a set of reliable tested and proven methods alongside step by step guidance This practical resource not only explores synthetic methodology but also reaction mechanisms and applications in medicinal chemistry and drug discovery The publication begins with an introductory chapter covering general principles and methodologies nomenclature and strategies of stereoselective synthesis Next it is divided into three parts Part One General Methods and Strategies Part Two Stereoselective Synthesis by Bond Formation including C C bond formation C H bond formation C O bond formation C N bond formation Other C heteroatom formation and other bond formation Part Three Methods of Analysis and Chiral Separation References in every chapter serve as a gateway to the literature in the field With this publication as their guide chemists involved in the stereoselective synthesis of drugs and natural products now have a single expertly edited source for all the methods they need [Ei Engineering Conference Index](#) ,1985 **Energy Research Abstracts** ,1993 [Advances in Nanostructured Materials for Multifunctional Applications: Synthesis, Characterization and Properties](#) Luz Stella Gomez-Villalba,2019-01-25 Special topic volume with invited peer reviewed papers only **The Encyclopedia of Vitamin E** Victor R. Preedy,Ronald Ross Watson,2007 Vitamin E is an important dietary constituent which helps in the defence against cellular damage The process of its absorption from food and its utilization by the body is an intricate series of reactions It is also used therapeutically in treating numerous diseases and conditions such as skin damage and the prevention of pathological lesions in major organs and has been shown to be an important factor in preventing heart disease and cancer Over 100 chapters from international contributors make this book the most comprehensive reference work in describing both the positive and negative effects and actions of Vitamin E Chapters are divided into subsections which cover nomenclature biochemical physical and chemical aspects of vitamin E related compounds dietary and nutritional influences and effects cocktails anti oxidants mixtures and novel analogues general physiological systems metabolism and metabolic stress brain neurological and optical systems reproductive systems fetus and infant musculo skeletal systems and exercise cardiovascular and pulmonary systems skin hepatic nephrotic and gastrointestinal systems immune and haematological systems and cancer *Anticancer Research* ,1981 **EPA Publications Bibliography** ,1984 *Government Reports Announcements* ,1975-10-03 **Government Reports Announcements & Index** ,1978 **EPA publications bibliography, 1977-1983** ,1983 **Acid Precipitation** ,1984 **Physics Briefs** ,1993 [Low-dielectric Constant Materials-synthesis and Applications in Microelectronics](#) Toh-Ming Lu,1995

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