

Chapter 10 - Chemical Quantities

10.1 The Mole: A Measurement of Matter:

Mole:

- A counting unit
- 6.02×10^{23} representative particles (atom, molecule, formula unit, ion) of a substance = **Avogadro's Number**

Converting # of particles to Moles:

$$\text{Moles} = \cancel{\text{Representative particles}} \left(\frac{1 \text{ mole}}{6.02 \times 10^{23} \cancel{\text{representative particles}}} \right)$$

Converting Moles to # of particles:

$$\text{Representative particles} = \cancel{\text{moles}} \left(\frac{6.02 \times 10^{23} \cancel{\text{representative particles}}}{1 \cancel{\text{mole}}} \right)$$

Molar Mass:

- Molar mass of an element = atomic mass (g)
Ex. 24.3 g = molar mass of Mg
- Molar mass = g/mol
- Molar mass of any element contains 1 mol or 6.02×10^{23} atoms of that element
- Molar mass of a compound = sum of the atomic masses of each element contained in 1 mole of the compound.

10.2 Mole-Mass and Mole-Volume Relationships:

Mole-Mass Relationship:

- Use molar mass (g/mol) of an element or compound to convert b/w the mass of a substance and the moles of a substance

$$\text{Mass (grams)} = \cancel{\# \text{ of moles}} \left(\frac{\cancel{\text{mass (grams)}}}{1 \cancel{\text{mole}}} \right)$$

$$\cancel{\# \text{ of Moles}} = \cancel{\text{Mass (grams)}} \left(\frac{1 \cancel{\text{mole}}}{\cancel{\text{mass (grams)}}} \right)$$

Chapter 10 Chemical Quantities Packet Answers

Christian Drosten



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