#### Chapter 10 Practice Problems & Test Review Chemical Quantities Answer Sheet

Show all work (no work, no credit) Answers in RED

How many moles is 750 billion atoms H in water? (1 billion = 1x10°)

H=1.25x10<sup>-12</sup> H moles H<sub>2</sub>O=6.25x10<sup>-13</sup>moles H<sub>2</sub>O

- How many carbon atoms are in a 9 mole gasoline? (gasoline = C<sub>8</sub>H<sub>18</sub>) 4.32×10<sup>23</sup> C atoms
- 3. What is the molar mass of Iron (III) Oxalate?

 $Fe_2(C_2O_4)_3 = 375.6 \text{ g/mol}$ 

How many moles are 500 grams Iron (III) Oxalate?
 1,33 mol Fe:(C:O4):

- How many grams is 1.78 moles Iron (II) Hydroxide?
   Fe(OH)<sub>2</sub> = 89.8 g/mol
   159.8 g Fe(OH)<sub>2</sub>
- Calculate the volume of 3.44 moles methane, CH<sub>4</sub> gas at STP: 77.06 liters CH<sub>4</sub>
- What is the density of Nitrogen Dioxide at STP?
   NO<sub>2</sub> 46g/mol 2.05 g/l NO<sub>2</sub>
- What is the volume of 3.24x10<sup>22</sup> molecules of Chlorine gas at STP?
   1.21 liters Cl<sub>2</sub>
- 9. How many H atoms and O atoms are in a 500 ml water bottle?

H atoms =  $3.33 \times 10^{25}$ O atoms =  $1.67 \times 10^{25}$ 

10. A dime is an alloy (a mixture of two solids) of 75% Copper (Cu) and 25% Nickel (Ni) and has a mass of 2.27 grams; how many Cu, Ni, and total atoms does a dime contain?

 $Cu = 1.61 \times 10^{22}$  total atoms =  $2.19 \times 10^{22}$  $Ni = 5.80 \times 10^{21}$ 

Given a container with 50-ml of Ethanol (C<sub>2</sub>H<sub>6</sub>O); the density of Ethanol is 0.789g/ml 11. Ethanol's molar mass 46 g/mol

- 12. Moles of ethanol 1,166 moles Ethanol
- 13. Carbon atoms 1.4x10<sup>24</sup> C atoms

# **Chapter 10 Chemical Quantities Practice Problems Answers**

**Ensheng Dong** 

**Chapter 10 Chemical Quantities Practice Problems Answers:** 

The book delves into Chapter 10 Chemical Quantities Practice Problems Answers. Chapter 10 Chemical Quantities Practice Problems Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chapter 10 Chemical Quantities Practice Problems Answers, encompassing both the fundamentals and more intricate discussions.

- 1. The book is structured into several chapters, namely:
  - Chapter 1: Introduction to Chapter 10 Chemical Quantities Practice Problems Answers
  - Chapter 2: Essential Elements of Chapter 10 Chemical Quantities Practice Problems Answers
  - Chapter 3: Chapter 10 Chemical Quantities Practice Problems Answers in Everyday Life
  - Chapter 4: Chapter 10 Chemical Quantities Practice Problems Answers in Specific Contexts
  - ∘ Chapter 5: Conclusion

Problems Answers.

- 2. In chapter 1, this book will provide an overview of Chapter 10 Chemical Quantities Practice Problems Answers. This chapter will explore what Chapter 10 Chemical Quantities Practice Problems Answers is, why Chapter 10 Chemical Quantities Practice Problems Answers is vital, and how to effectively learn about Chapter 10 Chemical Quantities Practice Problems Answers.
- 3. In chapter 2, the author will delve into the foundational concepts of Chapter 10 Chemical Quantities Practice Problems Answers. This chapter will elucidate the essential principles that need to be understood to grasp Chapter 10 Chemical Quantities Practice Problems Answers in its entirety.
- 4. In chapter 3, the author will examine the practical applications of Chapter 10 Chemical Quantities Practice Problems Answers in daily life. This chapter will showcase real-world examples of how Chapter 10 Chemical Quantities Practice Problems Answers can be effectively utilized in everyday scenarios.
- 5. In chapter 4, the author will scrutinize the relevance of Chapter 10 Chemical Quantities Practice Problems Answers in specific contexts. This chapter will explore how Chapter 10 Chemical Quantities Practice Problems Answers is applied in specialized fields, such as education, business, and technology.
- 6. In chapter 5, the author will draw a conclusion about Chapter 10 Chemical Quantities Practice Problems Answers. The final chapter will summarize the key points that have been discussed throughout the book.

  The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chapter 10 Chemical Quantities Practice

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