

## Determining Empirical Formulas

1) 75% carbon, 25% hydrogen	Answer:	$\text{CH}_4$
Work: $\text{C } 75\% \rightarrow 75\text{g C} \left  \frac{1 \text{ mol}}{12.01\text{g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$ $\text{H } 25\% \rightarrow 25\text{g H} \left  \frac{1 \text{ mol H}}{1.01\text{g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{CH}_4</math></div>		
2) 52.7% potassium, 47.3% chlorine	Answer:	$\text{KCl}$
Work: $\text{K } 52.7\% \rightarrow 52.7\text{g K} \left  \frac{1 \text{ mol K}}{39.10\text{g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$ $\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left  \frac{1 \text{ mol Cl}}{35.45\text{g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{KCl}</math></div>		
3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen	Answer:	$\text{AlPO}_4$
Work: $\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left  \frac{1 \text{ mol Al}}{26.98\text{g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$ $\text{P } 25.4\% \rightarrow 25.4\text{g P} \left  \frac{1 \text{ mol P}}{30.97\text{g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$ $\text{O } 52.5\% \rightarrow 52.5\text{g O} \left  \frac{1 \text{ mol O}}{16.00\text{g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{AlPO}_4</math></div>		
4) 13% magnesium, 87% bromine	Answer:	$\text{MgBr}_2$
Work: $13\text{g Mg} \left  \frac{1 \text{ mol Mg}}{24.31\text{g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$ $87\text{g Br} \left  \frac{1 \text{ mol Br}}{79.90\text{g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{MgBr}_2</math></div>		
5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen	Answer:	$\text{Na}_2\text{SO}_4$
Work: $32.4\text{g Na} \left  \frac{1 \text{ mol}}{22.99\text{g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$ $22.5\text{g S} \left  \frac{1 \text{ mol S}}{32.37\text{g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$ $45.1\text{g O} \left  \frac{1 \text{ mol O}}{16\text{g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{Na}_2\text{SO}_4</math></div>		

# Determining Empirical Formulas Answer Key

**Mr. Ajeet Singh**



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**Prentice Hall Chemistry**, 2000

**Chemistry Homework** Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table, writing formulas, balancing

equations stoichiometry problems and more      **Introduction to Chemical Principles** Howard Stephen Stoker,1983

**Introduction to Chemistry, Laboratory Manual** T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students      Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998      **Addison-Wesley Chemistry**,1990      **Holt Chemistry** Ralph Thomas Myers,2004      *Preparatory Chemistry* Howard Stephen Stoker,1985      **Basic Chemistry** William S. Seese,Guido H. Daub,1977      **Students Study Guide** Fredeen,Pearson,2003-07      **Alchemy Revisited** C. G. Vlassis,1976      Organic Chemistry G. Marc Loudon,1988

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