

CHEMICAL BONDING

BASICS

DOODLE NOTES

Bonding Basics

Small structure

Atoms which are directly joined are called **adjacent**.



Adjacent atoms are joined by the **chemical bond** around the **adjacent**.

Br

Three Main Types of Chemical

The type of bond is determined by the **electronegativity** difference between the two **atoms**.

ionic bond

Electronegativity is **ability** of an atom to attract electrons towards itself. In ionic bonding, **electrons** are **transferred** from one atom to another, resulting in the formation of **ions** which are held together by **electrostatic forces**.

Notes on:
Three Main
types of
Chemical
Bonds

Covale
Bond

Formed between two **non-metal** atoms. The **electrons** are **shared** between the two atoms. Both atoms count the **shared** electrons as part of their **valence** shell.

Metall
Bond

Only forms between **metal** atoms.

Electronegativity difference between two atoms is **greater than 1.7**.

Covalent bonding	Electron dot diagram
<p>Example: <chem>CO2</chem></p>	<p>Example: <chem>KCl</chem></p>
Metallic bonding	Electron dot diagram
<p>Example: <chem>Cu</chem></p>	<p>Example: <chem>NP</chem></p>

Bonding Basics Guided Notes

A Gutmann



Bonding Basics Guided Notes:

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The Standard & Poor's Guide to Personal Finance Tom Downey,2005 Provides you with step by step guidelines for making the right decisions on everything from insurance and investments to credit cards college and retirement savings plans and other important day to day money issues **Forbes** ,1923 Student Study Guide to Accompany Fundamentals of Risk and Insurance Vaughan,1978-08 **Fundamentals of Chemistry, Study Guide** James E. Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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