

## Practicing Electron Configuration

1. Answer the following questions

- i. How many p orbitals can there be in an energy level? \_\_\_\_\_
- ii. Which is the lowest energy level with an s orbital? \_\_\_\_\_
- iii. Which is the lowest energy level that can have a p orbital? \_\_\_\_\_
- iv. How many d orbitals can there be in an energy level? \_\_\_\_\_
- v. How many d electrons can there be in an energy level? \_\_\_\_\_

2. Identify the element with the following electron configuration

- i.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$  \_\_\_\_\_
- ii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$  \_\_\_\_\_
- iii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$  \_\_\_\_\_
- iv.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$  \_\_\_\_\_
- v.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^1$  \_\_\_\_\_

3. Which elements are represented by each of the following electron configurations?

- i.  $1s^2 2s^2 2p^5$  \_\_\_\_\_
- ii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^4$  \_\_\_\_\_
- iii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$  \_\_\_\_\_
- iv.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$  \_\_\_\_\_

4. Which element has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$ ? \_\_\_\_\_

5. What is the electronic configuration of aluminum, Al? \_\_\_\_\_

6. What is the noble gas configuration for the Pb atom? \_\_\_\_\_

7. When two electrons occupy the same orbital, they must have the same or opposite spin? \_\_\_\_\_

8. Explain why the following electron configuration is invalid:  $1s^2 2s^2 2p^6 3s^3 3d^5$

\_\_\_\_\_

# Election Configuration Worksheet And Lots More Answers

**Liying Dong**



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