

- Probably not, it's water ( $\text{H}_2\text{O}$ )
- SATP** (standard ambient temperature & pressure:  $25^\circ\text{C}$ , 100 kPa).  
**STP** (standard T&P:  $0^\circ\text{C}$ , 101.325 kPa)  
**Element**: cannot be broken down by chemical means (all same type of atom). **Compound**: broken down by chemical means (two or more different types of atoms)  
**Group**: vertical column on periodic table  
**Period**: horizontal (left/right) row  
**Metal**: element to the left of the staircase line.  
They tend to be solids at SATP, and conduct.  
**Nonmetals**: to the right of staircase line. Tend to be gases, non-conductors, and brittle as solid

- Thompson**: with the identification of electrons he proposed that negative electrons existed in a positive dough.  
**Rutherford**: famous gold foil experiment proved that an atom was mostly empty, with a dense positive nucleus orbited by electrons  
**Bohr**: added to Rutherford's model the idea of "shells". Evidence includes line spectra.
- Mendeleev ordered table according to atomic mass (today it's done by atomic number)
  - Atomic number = # of protons  
Mass number = # of protons + # of neutrons  
They are averages (of different isotopes)
  - 20, 37, 17, 17



11. Covalent = a, d    Ionic = b, c, e

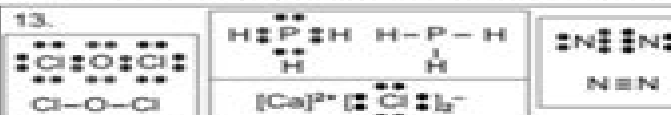
12. Lewis:



- In order from low to high boiling points:  
 $\text{H}_2$ : covalent ( $\Delta\text{EN}=0$ ),  $\text{CH}_4$ : covalent (0.4),  
 $\text{H}_2\text{O}$ : polar covalent (1.4),  $\text{LiF}$ : ionic (3.0)
- $\text{HCl}$ ,  $\text{Na}_2\text{O}$ ,  $\text{PCl}_3$ ,  $\text{Al}_2\text{O}_3$ ,  $\text{MgO}$
- a) copper(I) iodide, b)  $\text{HI}(\text{aq})$ , c) dinitrogen tetroxide, d) phosphorous acid, e)  $\text{PBr}_3$ , f)  $\text{Fe}_2\text{O}_3$ , g)  $\text{K}_2\text{N}$ , h)  $\text{H}_2\text{C}_2\text{O}_4$ , i) dichlorine heptoxide, j) hydrofluoric acid, k) nickel (II) sulfate hexahydrate, l) hydrogen sulfide
- combustion**:  $\text{AB} + \text{oxygen} \rightarrow \text{oxides of A \& B}$   
**synthesis**:  $\text{A} + \text{B} \rightarrow \text{C}$ , **decomposition**:  $\text{AB} \rightarrow \text{A} + \text{B}$   
**single displacement**:  $\text{A} + \text{BC} \rightarrow \text{AC} + \text{B}$   
**double displacement**:  $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$
- a) S, b) DD, c) S, d) D, e) SD, f) DD, g) S, h) SD

- I - Alkali metals, II - Alkaline earth metals, VII - halogens, VIII - noble gases, middle - transition elements/metals, bottom - inner transition elements (lanthanides, actinides)
- Democritus**: first to propose atom and the void based purely on logic  
**Aristotle**: his ideas were also based on thought. He was way off, but his ideas persisted for 2000 years  
**Dalton**: 1<sup>st</sup> to investigate structure of matter by experiment. His five postulates included the idea that all matter is made of atoms, each element has its own type of atoms, and atoms are rearranged in chemical reactions

- Size increases down a group (more shells), it decreases left to right as the # of protons increases, pulling outer electrons closer
- Ionization energy**: energy required to remove outer electron. It is high when atoms are small (high in group) with lots of protons (right in period). **Electron affinity**: the energy change when an electron adds to an atom. It is also high when atoms are small (high in group) with lots of protons (right in period).
- Electronegativity**: ability of atoms, when bonded, to attract electrons (essentially a numerical value for electron affinity). It follows the same trend as electron affinity for the same reason



- Ionic**: high melting/boiling points, soluble in polar solvents, conducts when dissolved in water but not as solid, brittle.  
**Covalent**: low melting/boiling points, soluble in non-polar solvents, doesn't conduct, soft.  
These differences are caused by the different strength of intermolecular forces in ionic versus covalent molecules.

- a)  $\text{Ca} + \text{CuSO}_4 \rightarrow \text{Cu} + \text{CaSO}_4$   
b)  $\text{FeCl}_2 + \text{Ag} \rightarrow \text{NR}$   
c)  $\text{H}_2\text{O} + \text{Ca} \rightarrow \text{H}_2 + \text{Ca}(\text{OH})_2$   
d)  $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2 + \text{Al}_2(\text{SO}_4)_3$   
e)  $\text{Na} + \text{Ni}_3(\text{PO}_4)_2 \rightarrow \text{Ni} + \text{Na}_3\text{PO}_4$   
f)  $\text{Au} + \text{HCl} \rightarrow \text{NR}$

# Bonding Of Atoms Chemistry Study Guide Answers

**Mary H. Bailey**



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## **Table of Contents Bonding Of Atoms Chemistry Study Guide Answers**

1. Understanding the eBook Bonding Of Atoms Chemistry Study Guide Answers
  - The Rise of Digital Reading Bonding Of Atoms Chemistry Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Bonding Of Atoms Chemistry Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bonding Of Atoms Chemistry Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Bonding Of Atoms Chemistry Study Guide Answers
  - Personalized Recommendations
  - Bonding Of Atoms Chemistry Study Guide Answers User Reviews and Ratings
  - Bonding Of Atoms Chemistry Study Guide Answers and Bestseller Lists
5. Accessing Bonding Of Atoms Chemistry Study Guide Answers Free and Paid eBooks

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  - Bonding Of Atoms Chemistry Study Guide Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
  - Bonding Of Atoms Chemistry Study Guide Answers Compatibility with Devices
  - Bonding Of Atoms Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Bonding Of Atoms Chemistry Study Guide Answers
  - Highlighting and Note-Taking Bonding Of Atoms Chemistry Study Guide Answers
  - Interactive Elements Bonding Of Atoms Chemistry Study Guide Answers
8. Staying Engaged with Bonding Of Atoms Chemistry Study Guide Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bonding Of Atoms Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Bonding Of Atoms Chemistry Study Guide Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bonding Of Atoms Chemistry Study Guide Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Bonding Of Atoms Chemistry Study Guide Answers
- Setting Reading Goals Bonding Of Atoms Chemistry Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Bonding Of Atoms Chemistry Study Guide Answers
- Fact-Checking eBook Content of Bonding Of Atoms Chemistry Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
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