

2 - Bonding & Molecular Structures

10 20 30 40 50 60 70 80 90 100

Table 4.1 - Some Average Single and Multiple Bond Strengths (kJ/mol)

| | H | C | N | O | F | Cl | Br | I | S | Se | Te |
|------|-----|-----|-----|-----|-----|-----|-----|-----|-----|-----|-----|
| H | 436 | | | | | | | | | | |
| C | | 413 | 347 | 336 | 485 | | | | | | |
| N | | | 391 | 305 | 460 | | | | | | |
| O | | | | 443 | 425 | | | | | | |
| F | | | | | 485 | | | | | | |
| Cl | | | | | | 328 | | | | | |
| Br | | | | | | | 276 | | | | |
| I | | | | | | | | 240 | | | |
| S | | | | | | | | | 265 | | |
| Se | | | | | | | | | | 226 | |
| Te | | | | | | | | | | | 208 |
| C-H | 413 | | | | | | | | | | |
| C-C | | 347 | | | | | | | | | |
| C-N | | | 391 | | | | | | | | |
| C-O | | | | 443 | | | | | | | |
| C-F | | | | | 485 | | | | | | |
| C-Cl | | | | | | 328 | | | | | |
| C-Br | | | | | | | 276 | | | | |
| C-I | | | | | | | | 240 | | | |
| C-S | | | | | | | | | 265 | | |
| C-Se | | | | | | | | | | 226 | |
| C-Te | | | | | | | | | | | 208 |

Multiple Bonds

| | | | |
|-------------|-----|-----|-----|
| C=C | 614 | C≡C | 839 |
| C=O | 799 | C≡N | 891 |
| C=N | 891 | N≡N | 945 |
| C=O and N=O | 607 | | |

Table 4.2 - Standard Enthalpies of Formation (kJ/mol)

| Compound | Standard Enthalpy of Formation | Compound | Standard Enthalpy of Formation |
|---------------------|--------------------------------|---------------------|--------------------------------|
| H ₂ (g) | 0 | H ₂ O(l) | -285.8 |
| O ₂ (g) | 0 | CO ₂ (g) | -393.5 |
| H ₂ O(g) | -241.8 | | |

1. Write the balanced chemical equation for the complete combustion of ethane (C₂H₆).



2. Draw structural formulae for each reactant and product species.



3. Calculate the standard enthalpy change for the reaction.

Calculate the standard enthalpy change for the reaction using the data in the table.



4. What is the standard enthalpy of formation of CO₂(g)?

5. Use the data in the table to calculate the standard enthalpy of formation of CO(g).

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Linus Pauling



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Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1967

Lectures On Chemical Bonding And Quantum Chemistry Sambhu N Datta, 2019-02-25 The concept of a chemical bond evolved from a variety of experimental observations It became useful to understand at times even predict the molecular structure reactivity and mechanism of chemical reactions Every aspect of the concept of bonding received a quantitative interpretation from the advent of quantum mechanics and its application to chemistry In Lectures on Chemical Bonding and Quantum Chemistry the reader will find a comprehensive discourse on the basic interpretation of the chemical bond as well as current understanding in terms of a dancing molecule that not only travels rotates and pulsates around an equilibrium molecular structure but also interacts and collides with other molecules thereby transferring linear and angular momentum characteristics and adjusting total energies One will also find a thorough survey of quantum mechanical methodologies for calculation of molecular characteristics in specific states and their changes under spectroscopic transitions tunneling electron and proton transfer phenomena and so on Guides to more advanced levels of theory are also provided

Chemical Bonding and Molecular Geometry Ronald James Gillespie, Paul L. A. Popelier, 2001 Provides an introduction to models and theories of chemical bonding and geometry as applied to the molecules of the main group elements This text also elucidates the relationships between these various models and theories It is useful for courses on chemical bonding in chemistry departments at the senior first year graduate level

Chemical Bonds Harry B. Gray, 2000

The Nature of the Chemical Bond and the Structure of Molecules and Crystals Linus Pauling, 1960 Thorough discussion of the various types of bonds their relative natures and the structure of molecules and crystals

Molecular Structure and Bonding Benjamin M. Gimarc, 1979

The Shape and Structure of Molecules Charles Alfred Coulson, 1982

Structure and Bonding Jack Barrett, 2002-05-09 Structure and Bonding covers introductory atomic and molecular theory as given in first and second year undergraduate courses at university level This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes Throughout the book theoretical concepts and experimental evidence are integrated An introductory chapter summarizes the principles on which the Periodic Table is established and describes the periodicity of various atomic properties which are relevant to chemical bonding Symmetry and group theory are introduced to serve as the basis of all molecular orbital treatments of molecules This basis is then applied to a variety of covalent molecules with discussions of bond lengths and angles and hence molecular shapes Extensive comparisons of valence bond theory and VSEPR theory with molecular orbital theory are included Metallic bonding is related to electrical conduction and semi conduction The energetics

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