#### AP Biology Study Guide 1

#### Water Properties:

Cohesion- Hydrogen bonding between LIKE molecules (ex. water with water)

Adhesion- Hydrogen bonding between UNLIKE molecules (ex. water with a leaf)

Capillary Action- the movement of a liquid through a narrow space as a result of the forces of cohesion, adhesion, and surface tension.

(ex. Moving ground water from wet areas of soil to dry ones.)

<u>Universal Solvent</u>- Water is referred to as a universal solvent because it dissolves more substances than any other liquid.

Hydration Shell- the sphere of water molecules around each dissolved ion.

Moderation of Temperature- water moderates air temperature by absorbing heat from the air that is warmer and releasing stored heat to the air that is cooler.

<u>Density</u>- water is LESS dense as a SOUD than a LIQUID (this is because of water's hydrogen bonds)

Polarity- water is a polar molecule because of the unequal distribution of electrons.

<u>High Specific Heat</u>- water has an unusually HIGH specific heat, this is how water is able to maintain its homeostasis and keep the ocean habitable.

#### Macrom olecules:

#### Carbohydrates

Elemental Makeup: Carbon, Hydrogen and Oxygen

Monomer: Monosaccharides (ex. Glucose & Fructose)

Polysaiccharides:

Cellulose: Strengthens cell walls
Starch: Stores glucose for energy (plants)
Glycogen: Stores glucose for energy (animals)
Chitin: Strengthens exoskeletons and fungal cell walls

Key Function: Fuel, Carbon source to be converted to other molecules or into polymers

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