

Sample Exercise 24.1 Identifying the Coordination Sphere of a Complex

Palladium(II) tends to form complexes with a coordination number of 4. One such compound was originally formulated as $\text{PdCl}_2 \cdot 3 \text{NH}_3$. **(a)** Suggest the appropriate coordination compound formulation for this compound. **(b)** Suppose an aqueous solution of the compound is treated with excess $\text{AgNO}_3(aq)$. How many moles of $\text{AgCl}(s)$ are formed per mole of $\text{PdCl}_2 \cdot 3 \text{NH}_3$?

Solution

Analyze: We are given the coordination number of Pd(II) and a chemical formula indicating that NH_3 and Cl^- are the potential ligands. We are asked to determine **(a)** what ligands are attached to Pd(II) in the compound and **(b)** how the compound behaves toward AgNO_3 in aqueous solution.

Plan: **(a)** Because of their charge, the Cl^- ions can be either in the coordination sphere, where they are bonded directly to the metal, or outside the coordination sphere, where they are bonded ionically to the complex. Because the NH_3 ligands are neutral, they must be in the coordination sphere. **(b)** The chlorides that are in the coordination sphere will not be precipitated as AgCl .

Solve:

(a) By analogy to the ammonia complexes of cobalt(III), we predict that the three NH_3 groups of $\text{PdCl}_2 \cdot 3 \text{NH}_3$ serve as ligands attached to the Pd(II) ion. The fourth ligand around Pd(II) is one of the chloride ions. The second chloride ion is not a ligand; it serves only as an anion in this ionic compound. We conclude that the correct formulation is $[\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{Cl}$.

(b) The chloride ion that serves as a ligand will not be precipitated as $\text{AgCl}(s)$ following the reaction with $\text{AgNO}_3(aq)$. Thus, only the single “free” Cl^- can react. We therefore expect to produce 1 mol of $\text{AgCl}(s)$ per mole of complex. The balanced equation is $[\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{Cl}(aq) + \text{AgNO}_3(aq) \rightarrow [\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{NO}_3(aq) + \text{AgCl}(s)$. This is a metathesis reaction (Section 4.2) in which one of the cations is the $[\text{Pd}(\text{NH}_3)_3\text{Cl}]^+$ complex ion.

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