

Thermochemistry Review - Answers

1. Using the data below, calculate the amount of energy required to melt a 16.2 g aluminum pop can. Assume that the aluminum is at an initial temperature of 20°C.

- Molar heat of fusion of Al(s) = 10.7 kJ/mol
- Melting point of Al(s) = 660°C
- Heat capacity of Al(s) = 0.900 J/g°C

$$\text{mol} = 16.2 \text{ g} \div 26.98154 \text{ g/mol} = 0.600 \text{ mol}$$

$$\Delta H = (20^\circ\text{C} \rightarrow 660^\circ\text{C}) + (\text{melt})$$

$$\Delta H = (mc\Delta T) + (n\Delta H_{\text{fus}})$$

$$\Delta H = (16.2 \text{ g})(0.900 \text{ J/g}^\circ\text{C})(640^\circ\text{C}) + (0.600 \text{ mol})(10.7 \text{ kJ/mol})$$

$$\Delta H = 9331.2 \text{ J} + 6.42 \text{ kJ}$$

$$\Delta H = 9.33 \text{ kJ} + 6.42 \text{ kJ}$$

$$\Delta H = \mathbf{15.8 \text{ kJ}}$$

2. Given the thermochemical equations:



Calculate the ΔH for the reaction: $\text{CH}_3\text{OH}(\text{l}) + \frac{3}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$



Ch Thermochemistry Answers

Claire Aldridge, Karl Lee



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Solutions Manual to Accompany Elements of Physical Chemistry C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book

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