

# Chapter 10 Chemical Quantities Review

## Part 1: Vocabulary

Label the following descriptions with the correct vocabulary term.

- |   |   |
|---|---|
| _____ mass  | _____ percent composition.                                |
| _____ empirical formula   | _____ empirical formula                                   |
| _____ molar mass  | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 <sup>24</sup>  |   |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition                                  |   |
| _____ the lowest whole-number ratio of the atoms of the elements in a compound _____ empirical formula                                  |   |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L   |   |
| _____ the smallest particle of a compound is a compound _____ molar mass  |   |
| _____ the number representing 6.02 $\times 10^{23}$ representative particles of a substance _____ Avogadro's number                     |   |
| _____ 1 mole = 6.02 $\times 10^{23}$ particles _____ empirical formula  |   |
| _____ the number representing the number of atoms of an element in a compound (2H <sub>2</sub> O _____ 2 hydrogen atoms, 1 oxygen atom) |   |

## Part 2: Section Concept Review

### 10.1 The Mole: A Measurement

**Completion** Use the completion strategy to strengthen knowledge of the terms and your understanding of the concepts involved with the terms of the definitions completed with a term (don't guess) or explain.

Measuring whole units of something, or mass, and its volume for a single quantity called the \_\_\_\_\_ mole. \_\_\_\_\_ Representative representative particles for 1 mole of a substance is \_\_\_\_\_ Avogadro's number \_\_\_\_\_

For the mass of 1 mole of a compound, scientists use together the \_\_\_\_\_ atomic masses \_\_\_\_\_ of the atoms making the compound. When we calculate the average for atoms (the weight the \_\_\_\_\_ molar mass \_\_\_\_\_ of the compound). These are \_\_\_\_\_ 6.02  $\times 10^{23}$  \_\_\_\_\_ representative particles the mass of one substance.

#### True or False

1 mole of each mass compound is 12 or 16 g.

True or False: 1 mole of substance contains 6.02  $\times 10^{23}$  atoms.

True or False: The representative particles of a ionic compound is the formula unit. **True**

True or False: 1 mole of H<sub>2</sub>O is composed of one mole of hydrogen and two atoms of oxygen.

True or False: One mole mass of nitrogen gas (N<sub>2</sub>) is \_\_\_\_\_ **28.02 g** **28.02 g**

#### Problems

# Chapter 10 Chemical Quantities Answer Key

**G Thomas**



## Chapter 10 Chemical Quantities Answer Key:

**Chemistry: Core Concepts, 3rd Edition** Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package

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**Introduction to Chemistry** Michael Mosher,Paul Kelter,2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry Meyer Brothers

Druggist ,1910 *Encyclopedia of Chemical Processing (Online)* Sunggyu Lee,2005-11-01 This second edition Encyclopedia supplies nearly 350 gold standard articles on the methods practices products and standards influencing the chemical industries It offers expertly written articles on technologies at the forefront of the field to maximize and enhance the research and production phases of current and emerging chemical manufacturing practices and techniques This collecting of information is of vital interest to chemical polymer electrical mechanical and civil engineers as well as chemists and chemical researchers A complete reconceptualization of the classic reference series the Encyclopedia of Chemical Processing and Design whose first volume published in 1976 this resource offers extensive A Z treatment of the subject in five simultaneously published volumes with comprehensive indexing of all five volumes in the back matter of each tome It includes material on the design of key unit operations involved with chemical processes the design unit operation and integration of reactors and separation systems process system peripherals such as pumps valves and controllers analytical techniques and equipment and pilot plant design and scale up criteria This reference contains well researched sections on automation equipment design and simulation reliability and maintenance separations technologies and energy and environmental issues Authoritative contributions cover chemical processing equipment engineered systems and laboratory apparatus currently utilized in the field It also presents expert overviews on key engineering science topics in property predictions measurements and analysis novel materials and devices and emerging chemical fields ALSO AVAILABLE ONLINE

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