

# Chapter 10 Chemical Quantities Review

## Part 1: Vocabulary

Label the following descriptions with the correct vocabulary term.

- |   |   |
|---|---|
| _____ mass  | _____ percent composition.                                |
| _____ empirical formula   | _____ empirical formula                                   |
| _____ molar mass  | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 <sup>23</sup>  |   |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition  |   |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula  |   |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L   |   |
| _____ the smallest particle of a compound is a compound _____ molar mass  |   |
| _____ the number representing 6.02 x 10 <sup>23</sup> representative particles of a substance _____ Avogadro's number                         |   |
| _____ 1 mole = 6.02 <sup>23</sup> particles _____ empirical formula   |   |
| _____ the number representing the relative mass of an element is atomic mass _____ 12.01 _____ number representing a substance's (STP, 1 atm) |   |

## Part 2: Section Concept Review

### 10.1 The Mole: A Measurement

**Completion** Use the completion strategy to strengthen knowledge of the terms and your understanding of the concepts involved with the terms of the definitions completed with a term (don't guess) or explain.

Measuring what is counting, or mass, and its volume for a single quantity called the \_\_\_\_\_ mole. \_\_\_\_\_ Representative representative particles for 1 mole of a substance is \_\_\_\_\_ Avogadro's number \_\_\_\_\_

It is the mass of \_\_\_\_\_ of a compound, which is used together with \_\_\_\_\_ percent mass \_\_\_\_\_ of the substance to find the percentage. Other two techniques for finding the mass (mole) of the \_\_\_\_\_ molar mass \_\_\_\_\_ of the compound. These are \_\_\_\_\_ 6.02 x 10<sup>23</sup> \_\_\_\_\_ representative particles the mass of one substance.

#### True or False

\_\_\_\_\_ 1 mole of any substance has a mass of 1 gram.

\_\_\_\_\_ 1 mole of substance contains 6.02 x 10<sup>23</sup> atoms.

\_\_\_\_\_ 1 mole of representative particles of a substance is the substance. **Incorrect**

\_\_\_\_\_ 1 mole of 10<sup>23</sup> is composed of one mole of carbon and four atoms of oxygen.

\_\_\_\_\_ 1 mole mass of nitrogen gas (N<sub>2</sub>) is **Incorrect** **nitrogen gas's molar mass**

#### Problems

# Chapter 10 Chemical Quantities Quiz Answer Key

**John Albert Sleicher**



## Chapter 10 Chemical Quantities Quiz Answer Key:

Chemical Surety United States. Department of the Army, 1995      *Foundations of Electronics* Russell L. Meade, 1999  
Learn electronics fundamentals for both DC and AC circuits from Ohms Law through series and parallel resonant circuits  
This highly acclaimed introduction to the world of electronics technology has been carefully updated to better provide technicians with a foundation in modern electronics needed to launch a career or pursue more advanced study Real world color codes and strategic highlighting are integrated with ample color charts photos schematics and diagrams for a solid understanding of circuit behavior that equips readers to progress to more complex topics with ease This edition features all new automated calculations for the formulas in the book on the accompanying CD as well as new information on admittance susceptance and more      **Fundamentals of Urine and Body Fluid Analysis - E-Book** Nancy A. Brunzel, 2012-03-23  
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**General, Organic, & Biological Chemistry** Janice Gorzynski Smith,2024-12 My goal in writing this text was to relate the fundamental concepts of general organic and biological chemistry to the world around us and in this way illustrate how chemistry explains many aspects of everyday life A key feature is the use of molecular art to illustrate and explain common phenomena we encounter every day Each topic is broken down into small chunks of information that are more manageable and easily learned Students are given enough detail to understand basic concepts such as how soap cleans away dirt and why trans fats are undesirable in the diet without being overwhelmed Introductory Chemistry Charles H. Corwin,2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking *Student's Guide to Introduction to Chemical Principles* by Edward I. Peters, 2d Ed Peter P. Berlow,1978 Holt Chemistry Ralph Thomas Myers,2004 **Meyer Brothers Druggist** ,1910 *Scientific American* ,1891 Monthly magazine devoted to topics of general scientific interest Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston,Denniston Katherine,2006 **Discover Science: Teacher's annotated edition** ,1991 Science content helps develop the skills needed to understand how science works learn new concepts solve problems and make decisions in today s technological society

**Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Children's Books in Print** R R

Bowker Publishing, Bowker, 1999-12  
**Newspaper** John Albert Sleicher, 1876

*English Mechanic and World of Science*, 1907  
The Economist, 1928

**Frank Leslie's Illustrated**

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