

Guide to Chapter 10. Liquids and Solids

Answers in red

We will spend three lecture days on this chapter.

Read the introductory paragraph to Chapter 10.

Read Section 10.1 Polar Covalent Bonds and Dipole Moments.

Learning Objective 1: Given the formula of a molecular compound determine if it is polar or non polar covalent.

Learning Objective 2: Know the definition of STP. Know the standard molar volume (22.4 L)

Do Problems 1 - 4 at the end of the section.

Do the following end-of-chapter problems: 22, 30, 36, 38

Read Section 10.2 Intermolecular Forces and Section 10.3 Some Properties of Liquids

Learning Objective 3: Given the formula of a compound, determine the major attractive force. (ionic = ion-ion and ion-dipole) and (molecular and covalent = dipole-dipole, hydrogen bonding, or London dispersion forces).

Learning Objective 4: Given a list of compound formulas rank their attractive forces in order of strength.

Learning Objective 5: Describe the relationship between size and strength of London dispersion forces.

Learning Objective 6: Describe the relationship between size and strength of London dispersion forces.

Do Problems 5 and 6 at the end of Section 10.2.

Do the following end-of-chapter problems: 32, 34, 40

Problem Club Question A. (a) Create a graph (general appearance only) with *MM* on the x-axis and boiling points on the y-axis for these similar compounds: NH_3 , PH_3 , AsH_3 , and SbH_3 . (b) Repeat for CH_4 , SiH_4 , GeH_4 , and SnH_4

Answer: boiling points (°C): (a) $\text{NH}_3 = -33$, $\text{PH}_3 = -87$, $\text{AsH}_3 = -55$, $\text{SbH}_3 = -17$. (b) $\text{CH}_4 = -164$, $\text{SiH}_4 = -111$, $\text{GeH}_4 = -89$, and $\text{SnH}_4 = -52$

Problem Club Question B. Why is iodine a solid?

*Answer: very substantial London-dispersion forces due to the large *MM**

Problem Club Question C. Circle the chemical with the highest melting point:

Answers in red

(a) F_2 or Cl_2

(b) F_2 or PCl_3

(c) Cl_2 or PCl_3

(d) NaCl or Cl_2

(e) HF or HBr

(f) $\text{CH}_3\text{CH}_2\text{CH}_3$ or $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$

Problem Club Question D. True or False:

Answers in red

T F Covalent bonds are weaker than London-dispersion forces

T F Dipole forces are stronger than London-dispersion forces

T F Only non-polar molecules have London-dispersion forces

T F CaS and KCl have similar *MMs*, however CaS should have a much larger mp.

T F The heat of vaporization increases with *MM* for the series $(\text{BH}_3)_n$

Chapter 10 Liquids And Solids Answers

J. J. Kipling



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Table of Contents Chapter 10 Liquids And Solids Answers

1. Understanding the eBook Chapter 10 Liquids And Solids Answers
 - The Rise of Digital Reading Chapter 10 Liquids And Solids Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Liquids And Solids Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Liquids And Solids Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Liquids And Solids Answers
 - Personalized Recommendations
 - Chapter 10 Liquids And Solids Answers User Reviews and Ratings
 - Chapter 10 Liquids And Solids Answers and Bestseller Lists

5. Accessing Chapter 10 Liquids And Solids Answers Free and Paid eBooks
 - Chapter 10 Liquids And Solids Answers Public Domain eBooks
 - Chapter 10 Liquids And Solids Answers eBook Subscription Services
 - Chapter 10 Liquids And Solids Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chapter 10 Liquids And Solids Answers Compatibility with Devices
 - Chapter 10 Liquids And Solids Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 10 Liquids And Solids Answers
 - Highlighting and Note-Taking Chapter 10 Liquids And Solids Answers
 - Interactive Elements Chapter 10 Liquids And Solids Answers
8. Staying Engaged with Chapter 10 Liquids And Solids Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Liquids And Solids Answers
9. Balancing eBooks and Physical Books Chapter 10 Liquids And Solids Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Liquids And Solids Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Liquids And Solids Answers
 - Setting Reading Goals Chapter 10 Liquids And Solids Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 Liquids And Solids Answers
 - Fact-Checking eBook Content of Chapter 10 Liquids And Solids Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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