

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

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Table of Contents Ch 8 Covalent Bonding Packet Answers

1. Understanding the eBook Ch 8 Covalent Bonding Packet Answers
 - The Rise of Digital Reading Ch 8 Covalent Bonding Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 8 Covalent Bonding Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 8 Covalent Bonding Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 8 Covalent Bonding Packet Answers

- Personalized Recommendations
- Ch 8 Covalent Bonding Packet Answers User Reviews and Ratings
- Ch 8 Covalent Bonding Packet Answers and Bestseller Lists
- 5. Accessing Ch 8 Covalent Bonding Packet Answers Free and Paid eBooks
 - Ch 8 Covalent Bonding Packet Answers Public Domain eBooks
 - Ch 8 Covalent Bonding Packet Answers eBook Subscription Services
 - Ch 8 Covalent Bonding Packet Answers Budget-Friendly Options
- 6. Navigating Ch 8 Covalent Bonding Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ch 8 Covalent Bonding Packet Answers Compatibility with Devices
 - Ch 8 Covalent Bonding Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 8 Covalent Bonding Packet Answers
 - Highlighting and Note-Taking Ch 8 Covalent Bonding Packet Answers
 - Interactive Elements Ch 8 Covalent Bonding Packet Answers
- 8. Staying Engaged with Ch 8 Covalent Bonding Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 8 Covalent Bonding Packet Answers
- 9. Balancing eBooks and Physical Books Ch 8 Covalent Bonding Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 8 Covalent Bonding Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 8 Covalent Bonding Packet Answers
 - Setting Reading Goals Ch 8 Covalent Bonding Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 8 Covalent Bonding Packet Answers

- Fact-Checking eBook Content of Ch 8 Covalent Bonding Packet Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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