Gas Laws Review Sheet

Definitions:

- Boyle's Law states that as the pressure of a gas increases the volume decreases
- Charles' Law states that as the temperature of a gas increases the volume also increases
- Dalton's Law of Partial Pressures states that the sum of the partial pressures of individual gases is equal to the total pressure in a container

Formulas:

- What is the formula for the combined gas law? P1V1 = P2V2
- 5. What is the formula for the ideal gas law?
 PV = nRT
- 6. What is the value for R in the ideal gas law?
- What does R stand for in the ideal gas law?
- the ideal gas law constant

 8. What does n stand for in the ideal gas law?
- the number of moles

Short Answer:

- 9. What happens to the volume of a gas if the temperature is increased? the volume increases
- 10. What happens to the volume of a gas if the temperature is decreased? the volume decreases
- 11. What happens to the volume of a gas if the pressure is increased? the volume decreases
- 12. What happens to the volume of a gas if the pressure is decreased? the volume increases
- 13. What happens to the pressure of a gas if the temperature is increased? the pressure increases
- 14. What happens to the pressure of a gas if the temperature is decreased? the pressure decreases

Calculations:

15. If 3.2 moles of a gas are at a temperature of 300 K and a volume of 83.8 L, what is the pressure of the gas?

$$P = \frac{nRT}{V}$$
 $P = \frac{3.2 \times 8.31 \times 300}{83.8}$ $P = 95 \text{ kPa}$

16. If 2.4 L of gas is at a pressure of 387 kPa, what would the pressure be if the volume was expanded to 29.8 L?

$$P_2 = P_1V_1$$
 $P_2 = 387 \times 2.4$ $P_2 = 31 \text{ kPa}$ V_2 29.8

 Argon occupies a volume of 28.3 L at a pressure of 50 kPa. Find the volume of argon when the pressure is increased to 310 kPa.

$$V_2 = P_1V_1$$
 $V_2 = \frac{50 \times 28.3}{310}$ $V_2 = 4.6 L$

Challenging Gas Laws Review All Concepts Answers

Siddappa N.Byrareddy

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